

Chemistry I (H) - Unit 02 (Pt. II - Nuclear Chemistry) Study Guide

1) Define, compare, and contrast **radiation** (pg. 122) to **radioactivity** (pg. 122).

2) Define AND give an example of a **radioisotope** (pg. 861).

3) Define **nuclear reactions**.

- Define **nuclear equations** (pg. 123/869) -
- Example:

4) Define and provide examples of the following forms of radioactive decay:

- **alpha decay** (pg. 867) -
 - **alpha particle** (pg. 123/862) -
 - Example Nuclear Equation:
- **beta decay** (pg. 867) -
 - **beta particle** (pg. 123/863) -
 - Example Nuclear Equation:
- **gamma radiation** (pg. 124) -
 - **gamma ray** (pg. 124/863) -
 - Example Nuclear Equation:
- Study the Table 24.2 from pg. 861 that we drew in our notes about the properties of decay (penetrating power, mass, etc.)

5) Define AND give an example of **transmutation** (pg. 865).

6) What do we mean by the words **emit** or **emanate**? Give an example.

7) Compare a **stable nucleus** to that of an **unstable nucleus**.

Nuclear Chemistry Study Guide And Practice Problems

Frank L. Wiseman

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