

Chemistry I (H) - Unit 02 (Pt. II - Nuclear Chemistry) Study Guide

- 1) Define, compare, and contrast radiation (pg. 122) to radioactivity (pg. 122).
- 2) Define AND give an example of a radioisotope (pg. 861).
- 3) Define nuclear reactions.
 - Define nuclear equations (pg. 123/869) -
 - Example:
- 4) Define and provide examples of the following forms of radioactive decay:
 - alpha decay (pg. 867) -
 - alpha particle (pg. 123/862) -
 - Example Nuclear Equation:
 - beta decay (pg. 867) -
 - beta particle (pg. 123/863) -
 - Example Nuclear Equation:
 - gamma radiation (pg. 124) -
 - gamma ray (pg. 124/863) -
 - Example Nuclear Equation:
 - Study the Table 24.2 from pg. 861 that we drew in our notes about the properties of decay (penetrating power, mass, etc.)
- 5) Define AND give an example of transmutation (pg. 865).
- 6) What do we mean by the words emit or emanate? Give an example.
- 7) Compare a *stable nucleus* to that of an *unstable nucleus*.

Nuclear Chemistry Study Guide And Practice Problems

Frank L. Wiseman



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Ionic and covalent compounds Chemical reactions Kinetics Acids and bases Electrochemistry Organic chemistry Nuclear chemistry Stoichiometry Thermodynamics

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