

Chemical Bonding

1. OCTET RULE (LEWIS-KOSSEL RULE)

Noble gases have 8 electrons in their outermost shell (complete octet) and outermost configuration is ns^2p^6 . Every atom has a tendency to complete its octet by losing or gaining or by sharing electron.

1.1 Exceptions of Octet Rule

(i) **Electron-Deficient Compounds** : Compounds having less than 8 electrons in the outermost shell of the central atom. e.g. BF_3 , BeCl_2 , AlCl_3 .

(ii) **Expansion of Octet** : The compounds having more than 8 electrons in the outermost shell of the central atom. e.g PCl_5 , SF_6 .

(iii) **Transition Metal Ions** : Transition metal ions have 9 to 18 electrons in their outermost shell. e.g. Cr^{+3} , Mn^{+2} , Cu^{+} .

(iv) H and Li atoms attain configuration like He by gaining and losing one electron, respectively. e.g. H^{+1} , Li^{+1} .

(v) **Odd Electron Bond** : Compounds having unpaired electron are also exceptions to the octet rule. The molecules having odd electrons are NO_2 , NO_3^- , ClO_4^- , O_2^{+1} etc.

Ex.1 Octet rule is not followed in:

$$\text{[1]} \text{ Ti}^{+3} \quad \text{[2]} \text{ Na}^{+} \quad \text{[3]} \text{ Ca}^{+2} \quad \text{[4]} \text{ K}^{+1} \quad \text{Ans. [1]}$$

Ques. Tl^{+3} has 9 electron in its outermost shell.

Ex.2 Which of the following compounds has the central atom with incomplete octet ?

[1] NH_3 [2] BCl_3 [3] CCl_4 [4] PCl_5 Ans. [2]

Sol. B has 6 electron in BCl_3 .

Ex.3 An element forms cation by losing electron. The configuration of the cation may be like that of :

[1] inert gas [2] pseudo inert gas [3] inert pair [4] all above Ans. [4]

Sol. Inert gas — 8 electron in outermost shell. Pseudo inert gas — 18 electron in outermost shell.

1.2 Inert Pair Configuration

The heavy elements of IIIA, IVA, and VA groups form cations having charge equal to group number and two less than group number, e.g. Ti^{+3} , In^{+3} , Sn^{+2} , Pb^{+2} . In these elements due to more penetration power of s orbitals, the atom generally does not lose e^- present in ns orbital, so only np electrons are lost by the atom. This effect is called "inert pair effect".

2. IONIC BOND

(i) Atoms get established by ion formation in formation of an ionic bond. One atom forms a cation by losing electrons and other forms anion by accepting electrons.

(ii) The octets of the atoms are completed by transfer of electrons.

(iii) Cations and anions are bonded by electrostatic force of attraction in ionic bond.

(iv) Ionic bond is formed between metals (cations) and nonmetals (anions).

(v) During the formation of an ionic bond, cation can attain one of the following configurations.

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Chemical Bonding for JEE Main & Advanced, NEET 2nd Edition Vaibhav Trivedi, 2019-10-10 The thoroughly revised updated 2nd edition of the book on Chemical Bonding is designed especially in accordance with latest competitive trends. The book has been updated with the past questions of NEET JEE Main JEE Advanced. A new chapter entitled Hydrolysis of Covalent Compounds has been added based on student's high demand. The salient features of the book are as follows: A moderately concise and compact book covering all topics from A to Z. Bent Rule with latest amendments and Drago's Rule. Physical properties of ionic covalent compounds with detailed explanation. Increasing and decreasing order of lattice energy, hydration energy, polarization and effect of these on physical properties has been done comparatively. Simple language to make it useful even to average and weak students. Logical and evolutionary approach in descriptions for better imagination and visualization. Large no of solved examples, illustrations and Objective type questions. Miscellaneous Practice Problems as final challenge.

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- i Extensive and well written theory
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About the Author: Shishir Mittal received his B.Tech in chemical engineering from IIT Bombay and is currently training Students for IIT JEE in Kota.

Super Course in Chemistry for the IIT-JEE: Physical Chemistry ,

Chemical Bonds Jeremy K. Burdett, 1997-05-21 Taking a question and answer format this tutorial explores the fundamental concepts of chemical bonding and is based on the lectures of Professor Jeremy Burdett.

Conceptual Chemistry Class XI Vol. I S K Jain, A book on Conceptual Chemistry

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Lectures On Chemical Bonding And Quantum Chemistry Sambhu N Datta, 2019-02-25

The concept of a chemical bond evolved from a variety of experimental observations. It became useful to understand at times

even predict the molecular structure reactivity and mechanism of chemical reactions Every aspect of the concept of bonding received a quantitative interpretation from the advent of quantum mechanics and its application to chemistry In Lectures on Chemical Bonding and Quantum Chemistry the reader will find a comprehensive discourse on the basic interpretation of the chemical bond as well as current understanding in terms of a dancing molecule that not only travels rotates and pulsates around an equilibrium molecular structure but also interacts and collides with other molecules thereby transferring linear and angular momentum characteristics and adjusting total energies One will also find a thorough survey of quantum mechanical methodologies for calculation of molecular characteristics in specific states and their changes under spectroscopic transitions tunneling electron and proton transfer phenomena and so on Guides to more advanced levels of theory are also provided

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Chemical Bonding Audrey L. Companion,1979 **Pearson IIT**

Foundation Chemistry Class 9 Trishna Knowledge Systems,2017 Pearson IIT Foundation Series one of the most reliable and comprehensive source of content for competitive readiness is now thoroughly updated and redesigned to make learning more effective and interesting for students The core objective of this series is to help aspiring students understand the fundamental concepts with clarity in turn helping them to master the art of problem solving Hence great care has been taken to present the concepts in a lucid manner with the help of neatly sketched illustrations and well thought out real life examples As a result this series is indispensable for any student who intends to crack high stakes examinations such as Joint Entrance Examination JEE National Talent Search Examination NTSE Olympiads Junior Senior International Kishore Vaigyanik Protsahan Yojana KVPY etc The series consists of 12 books spread across Physics Chemistry and Mathematics for

classes VII to X **Atoms, Chemical Bonds, and Bond Dissociation Energies** Sándor Fliszár, 1994 Chemical bonds their intrinsic energies in ground state molecules and the energies required for their actual cleavage are the subject of this book. The theory modelled after a description of valence electrons in isolated atoms explains how intrinsic bond energies depend on the amount of electronic charge carried by the bond forming atoms. It also explains how bond dissociation depends on these charges. While this theory vividly explains thermochemical stability future research could benefit from a better understanding of bond dissociation if we learn how the environment of a molecule affects its charges we also learn how it modifies bond dissociation in that molecule. This essay is aimed at theoretical and physical organic chemists who are looking for new perspectives to old problems.

Chemical Bonding at Surfaces and Interfaces Anders Nilsson, Lars G.M.

Pettersson, Jens Norskov, 2011-08-11 Molecular surface science has made enormous progress in the past 30 years. The development can be characterized by a revolution in fundamental knowledge obtained from simple model systems and by an explosion in the number of experimental techniques. The last 10 years has seen an equally rapid development of quantum mechanical modeling of surface processes using Density Functional Theory (DFT). Chemical Bonding at Surfaces and Interfaces focuses on phenomena and concepts rather than on experimental or theoretical techniques. The aim is to provide the common basis for describing the interaction of atoms and molecules with surfaces and this to be used very broadly in science and technology. The book begins with an overview of structural information on surface adsorbates and discusses the structure of a number of important chemisorption systems. Chapter 2 describes in detail the chemical bond between atoms or molecules and a metal surface in the observed surface structures. A detailed description of experimental information on the dynamics of bond formation and bond breaking at surfaces make up Chapter 3. Followed by an in depth analysis of aspects of heterogeneous catalysis based on the d band model. In Chapter 5 adsorption and chemistry on the enormously important Si and Ge semiconductor surfaces are covered. In the remaining two Chapters the book moves on from solid gas interfaces and looks at solid liquid interface processes. In the final chapter an overview is given of the environmentally important chemical processes occurring on mineral and oxide surfaces in contact with water and electrolytes. Gives examples of how modern theoretical DFT techniques can be used to design heterogeneous catalysts. This book suits the rapid introduction of methods and concepts from surface science into a broad range of scientific disciplines where the interaction between a solid and the surrounding gas or liquid phase is an essential component. Shows how insight into chemical bonding at surfaces can be applied to a range of scientific problems in heterogeneous catalysis, electrochemistry, environmental science and semiconductor processing. Provides both the fundamental perspective and an overview of chemical bonding in terms of structure, electronic structure and dynamics of bond rearrangements at surfaces.

Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1963 Structure of Molecules and the Chemical Bond Y. K. Syrkin, M. E.

Dyatkina, 1950 **Chemical Bond In Inorganic Chemistry**, I. David Brown, 2007-10-29 Chemical Bonding Mark J.

Winter,2016

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