

Background

- 1. Study of this \_\_\_\_\_ multi \_\_\_\_\_ relationship is a \_\_\_\_\_ method \_\_\_\_\_
- 2. Research on \_\_\_\_\_ relationship \_\_\_\_\_ equated \_\_\_\_\_

2.  $40g = 40r + 4 \cdot 20g$

2. Magnesium nitride reacts with a single oxygen when the products of synthesis are at magnesium oxide

The \_\_\_\_\_ is a \_\_\_\_\_ reaction \_\_\_\_\_ give the \_\_\_\_\_ \_\_\_\_\_ from the \_\_\_\_\_ reaction in the \_\_\_\_\_

3. Problem

When aluminum aluminum reacts with aluminum oxide, aluminum chloride is produced



moles reaction



2. If you start with 4 moles of  $Al$ , how many moles of  $AlCl_3$  will be produced?

2 moles  $Al$  (moles  $AlCl_3$ ) and  $3g = 4$  and  $3g$

# Note Taking Guide Episode 801 Stoichiometry

**S Marginson**



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