

Today's CD rates



Cd Rates Update

Jessica J Manson

Cd Rates Update:

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2 $\text{CO}_2 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g}$ $\Delta h = 520 \text{ kJ}$
2 $\text{C}_2\text{H}_2 \text{ g} \rightarrow 2 \text{ H}_2 \text{ g}$ $\Delta h = 226 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + 2 \text{ H}_2 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2\text{O} \text{ g}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + 7/2 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ g}$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + 2 \text{ H}_2 \text{ g}$ $\Delta h = 1560 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6 \text{ g} + 7/2 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ g}$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2 \text{ g} + 1/2 \text{ O}_2 \text{ g} \rightarrow \text{H}_2\text{O} \text{ g}$ $\Delta h = 285 \text{ kJ}$ mole

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 $\text{C}_2\text{H}_4 \text{ g} \rightarrow \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g}$ 2 $\text{H}_2\text{O} \text{ l} \text{ h}$ 1411 kj
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} \rightarrow \text{H}_2\text{O} \text{ l} \text{ h}$ 1560 kj

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web thus for reaction 9 2 in which 1 mole of H_3O and 1 mole of OH combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kj}$ as shown note that the enthalpy change Δh is positive if heat is absorbed that is if Δh is negative

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH} \text{ s} \rightarrow \text{Na} \text{ aq} + \text{OH} \text{ aq}$ $\Delta h = -12 \text{ kJ}$
2 $\text{NaOH} \text{ s} \rightarrow \text{H} \text{ aq} + \text{Cl} \text{ aq}$ $\Delta h = -20 \text{ kJ}$
3 $\text{Na} \text{ aq} + \text{Cl} \text{ aq} \rightarrow \text{NaCl} \text{ aq}$ $\Delta h = -5 \text{ kJ}$