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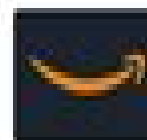
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web name the following molecular compounds  $\text{sc}_2$   $\text{n}_2\text{o}_4$   $\text{p}_4\text{o}_{10}$   $\text{pf}_5$   $\text{sc}_2$  sulfur dichloride  $\text{n}_2\text{o}_4$  dinitrogen tetroxide the a of tetra is dropped before the o of oxide  $\text{p}_4\text{o}_{10}$  tetraphosphorous decoxide the a of deca is dropped before the o of oxide  $\text{pf}_5$  phosphorous pentafluoride

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web one of the ions produced is always h which immediately combines with a water molecule to form the hydronium ion  $\text{h}_3\text{o}^+$  the  $\text{h}^+$  in is what defines the acidic properties of a substance because of the special classification acids have a naming system different from ionic or other molecular covalent compounds

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web acids are covalently bonded molecules but when they are put into water they produce ions one of the ions produced is always h which immediately combines with a water molecule to form the hydronium ion  $\text{h}_3\text{o}^+$  the  $\text{h}_3\text{o}^+$  ion is what defines the



acidic properties of a substance

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web examine the pairs of ternary acids in model 2 that contain sulfur phosphorus and nitrogen each pair has one acid that ends in ic and another that ends in ous these endings are related to the name of the polyatomic anion found in the acid ate or ite

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web the acids lose a hydrogen ion for each acid base reaction in model 2 describe the role of the brønsted lowry base in the proton h ion transfer that occurs the base gains or accepts a hydrogen ion proton as you saw in last week's pogil all arrhenius bases in model 1 have an oh ion in their chemical formulas

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web acid base reactions can be understood in terms of the transfer of hydrogen ion and the concentration of hydrogen ion in a reaction mixture measured as ph is often an important factor in how reactions occur learning objectives understand the brønsted lowry theory of acids and bases understand the concepts of conjugate acid base pairs

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web one of the ions produced is always h which immediately combines with a water molecule to form the hydronium ion h<sub>3</sub>o the h<sub>3</sub>o ion is what defines the acidic properties of a substance because of their special classification acids have a naming system different from ionic or other molecular covalent compounds

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web oct 7 2015 we will address the question what makes something an acid strong and strong how is it named  
 particledescriptiontable 1 particle descriptions strong and strong names of various acids particles not drawn to scale cl clbr  
 brf h hhhhhf hcl hbr hhfcl hbr hfformulahcl hbr hfnamehydrochloric acid

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web the molecular or empirical formula acids have their own set of naming rules cf sec 2 8 which you must also know you also need to be able to recognize and name a few organic compounds such as simple alkanes and alcohols see sec 2 9 key questions 8 in the chemical formula and name which element is given first a metal or nonmetal 9

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web each pair has one acid that ends in ic and another that ends in ous these endings are related to the name of the polyatomic anion found in the acid ate or ite complete the statements below with the correct acid name ending polyatomic anion ending is ate acid name ending is if