

CHEMISTRY FORM THREE NOTES

CHEMISTRY OF CARBON

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A: CARBON

Carbon is an element in Group IV (Group 4) of the Periodic table. It has atomic number 6 and electronic configuration $2:4$ and thus has four valence electrons (tetravalent). It does not easily ionize but forms strong covalent bonds with other elements including itself.

(a) Occurrence

Carbon mainly naturally occurs as:

- (i) allotropes of carbon i.e. graphite, diamond and fullerenes.
- (ii) amorphous carbon in coal, peat, charcoal and coke.
- (iii) carbon(IV) oxide gas accounting 0.03% by volume of normal air in the atmosphere.

(b) Allotropes of Carbon

Carbon naturally occurs in two main crystalline allotropic forms, carbon-graphite and carbon-diamond

Carbon-diamond	Carbon-graphite
Shiny crystalline solid	Black/dull crystalline solid
Has a very high melting/boiling point because it has a very closely packed giant tetrahedral structure joined by strong covalent bonds	Has a high melting/boiling point because it has a very closely packed giant hexagonal planar structure joined by strong covalent bonds
Has very high density (Hardest known natural substance)	Soft
Abrasive	Slippery
Poor electrical conductor because it has no free delocalized electrons	Good electrical conductor because it has free 4^{th} valency delocalized electrons
Is used in making Jewels, drilling and cutting metals	Used in making Lead-pencils, electrodes in batteries and as a lubricant
Has giant tetrahedral structure	Has giant hexagonal planar structure

c) Properties of Carbon

(i) Physical properties of carbon

Carbon occurs widely and naturally as a black solid

It is insoluble in water but soluble in carbon disulphide and organic solvents.

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The APPGCET Chemical Sciences PDF Andhra Pradesh Post Graduate Common Entrance Test Chemical Sciences Subject PDF eBook Covers Objective Questions From Various Competitive Exams With Answers Analytical Chemistry of Synthetic Colorants A.T. Peters,J. Szep,2013-03-09 More than one and a half decades have passed since the last book was published describing developments in the analytical chemistry of synthetic colorants In the intervening period the scope and technical capabilities of instrumentation for analysing dyes and pigments has significantly expanded It is now possible to rapidly resolve a number of problems whose solutions were previously either unattainable or very difficult to achieve For instance the unambiguous assignment of all the signals in the proton NMR spectrum of a trisazo direct dye and the confirmation of the molecular weight of involatile and in particular sulphonated dyes without derivatisation are now routine analytical techniques in many laboratories today In addition it is now possible to record the NMR spectrum of a dye molecule on less than 1 mg of material and we are no longer limited to solution spectra since solid samples can now be routinely analysed in NMR experiments Whilst not attempting to be all encompassing this volume is intended to bridge the gap between what was covered in the earlier work edited by Professor Venkataraman and the developments which have since ensued in some key areas It provides important updates in X ray crystallography proton NMR IR spectroscopy and mass spectrometry and additionally covers topics such as ESR micro spectrophotometry and emission spectroscopy New Scientist ,2007

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