

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that represents the total energy of a system, including internal energy and the energy required to displace the environment.

$$H = U + PV$$

 where U is internal energy, P is pressure, and V is volume.
 - **Enthalpy Change** (ΔH) is the change in enthalpy during a process.

$$\Delta H = H_{\text{final}} - H_{\text{initial}}$$

 - **Exothermic Process**: A process that releases heat to the surroundings, resulting in a negative ΔH .
 - **Endothermic Process**: A process that absorbs heat from the surroundings, resulting in a positive ΔH .



- **Free Energy** (G) is a thermodynamic property that represents the maximum useful work that can be extracted from a system at constant temperature and pressure.

$$G = H - TS$$

 where H is enthalpy, T is temperature, and S is entropy.
 - **Free Energy Change** (ΔG) is the change in free energy during a process.

$$\Delta G = \Delta H - T\Delta S$$

 - **Spontaneous Process**: A process that occurs without the need for external energy input, resulting in a negative ΔG .
 - **Non-spontaneous Process**: A process that requires external energy input to occur, resulting in a positive ΔG .



- **Chemical Equilibrium** is a state in which the concentrations of reactants and products remain constant over time.

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

 where K is the equilibrium constant, $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products, and a , b , c , and d are the stoichiometric coefficients.
 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift in a direction that opposes the change.
 - **Reaction Quotient** (Q) is a measure of the relative amounts of reactants and products in a system at any given time.

$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

 - **Equilibrium Constant** (K) is a measure of the relative amounts of reactants and products in a system at equilibrium.

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

- **Reaction Rate** is the change in concentration of a reactant or product over time.

$$\text{Rate} = -\frac{1}{a} \frac{d[A]}{dt} = \frac{1}{c} \frac{d[C]}{dt}$$

 where Rate is the reaction rate, $[A]$ and $[C]$ are the concentrations of the reactant and product, and a and c are the stoichiometric coefficients.
 - **Rate Law** is an equation that relates the reaction rate to the concentrations of the reactants.

$$\text{Rate} = k[A]^m[B]^n$$

 where k is the rate constant, $[A]$ and $[B]$ are the concentrations of the reactants, and m and n are the reaction orders.

- **Activation Energy** (E_a) is the minimum energy required for a reaction to occur.

$$E_a = E_{\text{transition state}} - E_{\text{reactants}}$$

 where E_a is the activation energy, $E_{\text{transition state}}$ is the energy of the transition state, and $E_{\text{reactants}}$ is the energy of the reactants.
 - **Arrhenius Equation** is an equation that relates the rate constant (k) to the activation energy (E_a) and the temperature (T).

$$\ln k = -\frac{E_a}{RT} + \ln A$$

 where k is the rate constant, E_a is the activation energy, R is the gas constant, T is the temperature, and A is the pre-exponential factor.



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