

CARD 97

The Rate Equation

The rate of a reaction is measured by observing the rate of change of quantities of reactants or products, over time.

The Rate Equation
For a general reaction $A + B \rightarrow C$, where both A and B affect rate:
Rate (mol dm⁻³ s⁻¹) = $k[A]^m[B]^n$

k , the rate constant, is a proportionality constant, which is equal to the rate of the reaction when the concentration of all species in the rate equation is 1 mol dm⁻³. k is independent of the concentrations of the reactants.

m and n are the orders of the reaction with respect to the concentrations of A and B, respectively.

For a zero order: rate is $[A]$

SPEC 5.13a,b

CARD 97

The Rate Equation

The table below shows a series of initial rate experiments used to investigate the reaction between gases A, B and C. Determine the rate equation.

Experiment	Initial [A] / mol dm ⁻³	Initial [B] / mol dm ⁻³	Initial [C] / mol dm ⁻³	Initial rate / mol dm ⁻³ s ⁻¹
1	0.1	0.1	0.1	0.1
2	0.2	0.1	0.1	0.2
3	0.1	0.2	0.1	0.2
4	0.1	0.1	0.2	0.1

Compare 1 & 2: [A] doubled in exp 2.
Compare 2 & 3: [B] doubled in exp 3.
Compare 1 & 4: We first take into account the effect of [A] on the rate. We can use the affected rate constant to find the order of [A].
 $2.00 \times 10^{-4} = k \cdot 0.2^m \cdot 0.1^n$
 $0.1 = k \cdot 0.1^m \cdot 0.1^n$
So rate is 2.00×10^{-4} mol dm⁻³ s⁻¹. **2 marks**

SPEC 5.13a,b

Finding k

SPEC 5.13c

Following reaction is second order with respect to NO₂, and zero order with respect to CO.
 $2\text{NO}_2 + \text{CO} \rightarrow \text{N}_2\text{O}_4 + \text{CO}_2$

Write equation for this reaction.

1/2 (2 mark) (1 mark)

Write equation for this reaction.

1/2 (2 mark) (1 mark)

Write equation for this reaction.

1/2 (2 mark) (1 mark)

Write equation for this reaction.

1/2 (2 mark) (1 mark)

CARD 106

Structure of Benzene

SPEC 6.13a,b,f

Benzene consists of a planar hexagon of C atoms, each bonded to 1 H atom.

Each carbon atom in benzene uses 3 of its outer 4 electrons in bonding to 2 other carbon atoms and to 1 hydrogen atom.

Each carbon atom also has one electron in a p-orbital, which is at a right angles to the plane of the ring.

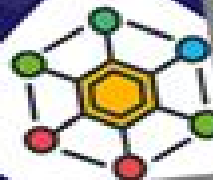


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