

The Rate Equation

CARD
9.1

SPHC
9.1.1 to 9.1.3

The rate of a reaction is measured by observing the rate of change of quantities of reactants or products over time.

The Rate Equation

For a general reaction $A + B \rightarrow C$, where both A and B affect rate:

$$\text{Rate} = k[A]^x[B]^y$$

• the rate constant, k , is a proportionality constant, which is equal to the rate of the reaction when the concentration of all species in the rate equation is 1 mol dm⁻³, with respect to the

The Rate Equation

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Experiment	Initial (A) / mol dm ⁻³	Initial (B) / mol dm ⁻³	Initial rate / mol dm ⁻³ s ⁻¹
1	0.1	0.1	0.001
2	0.2	0.1	0.004
3	0.1	0.2	0.002
4	0.2	0.2	0.008

Compare 1 & 2: $[\text{A}]$ doubled in exp 2, $[\text{B}]$ doubled in exp 2. Compare 2 & 3: $[\text{B}]$ doubled in exp 3. Compare 1 & 4: we first take into account the affected rate constant can use the affected rate constant to find the order of B. $k = \text{Rate} / [\text{A}]^1 [\text{B}]^1$ $0.001 = 10^{-3} \times 1 \cdot 5^1 \times 0.1 = 1.5 \times 10^{-3}$ $k = 1.5 \times 10^{-3}$ (1 mark) Rate order of B = 1.

Overall reaction is second order with respect to NO_2 and zero order with respect to CO_2 .

Rate equation for this reaction:

$\text{Rate} = k[\text{NO}_2]^2$ (1 mark)

Half order of the reaction:

$\text{Rate} = k[\text{NO}_2]$ (1 mark)

Rate of the rate constant, k , for this reaction:

$k = 1.5 \times 10^{-3} \text{ mol}^{-2} \text{ dm}^3 \text{ mol}^{-1} \text{ s}^{-1}$ (1 mark)

Structure of Benzene

SPHC
9.1.1 to 9.1.3

Benzene consists of a planar hexagon of C atoms, each bonded to 3 atoms. carbon atoms are delocalising where 3 of its electrons are delocalising in delocalising to 2 other carbon atoms and to 1 hydrogen atom.

Hydrogen atoms which form other delocalised bonds in the ring are delocalised in the ring.



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1. Understanding the eBook Ocr As Chemistry Revision Guide
 - The Rise of Digital Reading Ocr As Chemistry Revision Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr As Chemistry Revision Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr As Chemistry Revision Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr As Chemistry Revision Guide
 - Personalized Recommendations
 - Ocr As Chemistry Revision Guide User Reviews and Ratings
 - Ocr As Chemistry Revision Guide and Bestseller Lists

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 - Ocr As Chemistry Revision Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr As Chemistry Revision Guide
 - Highlighting and Note-Taking Ocr As Chemistry Revision Guide
 - Interactive Elements Ocr As Chemistry Revision Guide
8. Staying Engaged with Ocr As Chemistry Revision Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr As Chemistry Revision Guide
9. Balancing eBooks and Physical Books Ocr As Chemistry Revision Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr As Chemistry Revision Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr As Chemistry Revision Guide
 - Setting Reading Goals Ocr As Chemistry Revision Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr As Chemistry Revision Guide
 - Fact-Checking eBook Content of Ocr As Chemistry Revision Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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