

# How to Name Chemical Compounds

Is it ionic or covalent?  
or does it contain a polyatomic ion followed by H<sup>+</sup>?

## Ionic compound

[Cation's name] [Anion's name]

**metals**

- same as element name  
ex. Al<sup>3+</sup>: aluminum
- if transition metal<sup>+</sup>, include charge in Roman numeral  
ex. Cu<sup>2+</sup>: copper (II)

**poly-anions**

- just use the name!  
ex. NH<sub>4</sub><sup>+</sup>: ammonium
- just use H<sub>3</sub>O<sup>+</sup>  
ex. SO<sub>4</sub><sup>2-</sup>: sulfate

\* exceptions (no Roman numeral):

silver ion: silver Ag<sup>+</sup>

gold ion: gold(3+) Au<sup>3+</sup>

cadmium ion: cadmium Cd<sup>2+</sup>

## Molecular compound

[H]-[1<sup>st</sup> element name] [H]-[2<sup>nd</sup> element root]-"ide"

**Acids**  
root and  
naming

↑ skip if "mono-"

↑ skip if 1<sup>st</sup> element is H

**Stem/Prefixes**

- 1: mono-
- 2: di-
- 3: tri-
- 4: tetra-
- 5: penta-
- 6: hexa-
- 7: hepta-
- 8: octa-
- 9: nona-
- 10: deca-



## Acids

Drop all H<sup>+</sup>, and look at the name of the remaining anion.

Acid	Anion	Acid	Anion
HCl	chloride	H <sub>2</sub> SO <sub>4</sub>	sulfate
H <sub>3</sub> PO <sub>4</sub>	phosphate	H <sub>2</sub> SO <sub>3</sub>	thiosulfate

# Naming Molecular Compounds Study Guide

**Susan McMurry**

## **Naming Molecular Compounds Study Guide:**

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