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新實用漢語課本

3

NEW  
PRACTICAL  
CHINESE  
READER

WORKBOOK

綜合練習冊 (第3冊)

綜合練習冊



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# New Practical Chinese Reader Workbook Vol 3

**Arturo Cuomo**



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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{H}^+ + \text{SO}_2 \rightarrow \text{Mn}^{2+} + \text{H}_2\text{O} + \text{HSO}_4^-$  what will be the stoichiometric coefficient of  $\text{HSO}_4^-$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^-$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{AuCN} + 2\text{e}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_2 + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3^{2-} + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te}$  10  $\text{ReO}_4^- + \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$  as  $\text{NO}_3^- + \text{IO}_3^- + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$  and  $2\text{H}^+$   $\text{C}_6\text{H}_6\text{O}_6$  and  $2\text{H}^+$   $2\text{H}^+$  and  $2\text{H}^+$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $2\text{H}^+$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $2\text{H}^+$  b  $2\text{H}^+$  and  $\text{C}_6\text{H}_8\text{O}_6$  c  $\text{C}_6\text{H}_8\text{O}_6$  and  $2\text{H}^+$  d  $2\text{H}^+$  and  $\text{C}_6\text{H}_8\text{O}_6$

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