

## REDOX REACTIONS.

### NOTES...

#### Intro

Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

#### Oxidation

##### (i) Loss of electron

- Example:  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$

##### - Sub-points:

- Loss of hydrogen
- Gain of oxygen
- Gain of **electronegative** element OR loss of electropositive element

#### Reduction

##### i) Gain of electron

- Example:  $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$

##### - Sub-points:

- Gain of hydrogen
- Loss of oxygen
- Loss of electronegative element OR gain of electropositive element

#### Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

### **REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)**

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example:  $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$

# Mcgraw Hill Oxidation Reduction Study Guide

**RJ Shavelson**

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