

## CHEMISTRY FORM THREE NOTES

### CHEMISTRY OF CARBON

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##### A: CARBON

Carbon is an element in Group IV (Group 4) of the Periodic table. It has atomic number 6 and electronic configuration  $2:4$  and thus has four valence electrons (tetravalent). It does not easily ionize but forms strong covalent bonds with other elements including itself.

##### (a) Occurrence

Carbon mainly naturally occurs as:

- (i) allotropes of carbon i.e. graphite, diamond and fullerenes.
- (ii) amorphous carbon in coal, peat, charcoal and coke.
- (iii) carbon(IV) oxide gas accounting 0.03% by volume of normal air in the atmosphere.

##### (b) Allotropes of Carbon

Carbon naturally occurs in two main crystalline allotropic forms, carbon-graphite and carbon-diamond

Carbon-diamond	Carbon-graphite
Shiny crystalline solid	Black/dull crystalline solid
Has a very high melting/boiling point because it has a very closely packed giant tetrahedral structure joined by strong covalent bonds	Has a high melting/boiling point because it has a very closely packed giant hexagonal planar structure joined by strong covalent bonds
Has very high density (Hardest known natural substance)	Soft
Abrasive	Slippery
Poor electrical conductor because it has no free delocalized electrons	Good electrical conductor because it has free $4^{\text{th}}$ valency delocalized electrons
Is used in making Jewels, drilling and cutting metals	Used in making Lead-pencils, electrodes in batteries and as a lubricant
Has giant tetrahedral structure	Has giant hexagonal planar structure

##### c) Properties of Carbon

##### (i) Physical properties of carbon

Carbon occurs widely and naturally as a black solid

It is insoluble in water but soluble in carbon disulphide and organic solvents.

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