

CHAPTER 6**STUDY GUIDE****Ionic Compounds and Metals****Lesson 6.1 Ion Formation**

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) **Chemical Bond**.

Such an attachment may form by the attraction of the positively charged

(2) **Nucleus** of one atom for the negatively charged

(3) **Electron** of another atom, or by the attraction of charged atoms, which are called (4) **Ions**. The attractions may also involve

(5) **Valence** electrons, which are the electrons in the outermost

(6) **Energy level**. The (7) **Noble Gas** are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) **Octet**. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) **Pseudo-Noble gas formations**.

For each statement below, write *true* or *false*.

False 10. A positively charged ion is called an anion.

True 11. Elements in group 1 lose their one valence electron, forming an ion with a $1+$ charge.

False 12. Elements tend to react so that they acquire the electron structure of a halogen.

True 13. A sodium atom tends to lose one electron when it reacts.

True 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.

False 15. A Cl^- ion is an example of a cation.

True 16. The ending *-ide* is used to designate an anion.

False 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

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Karen Timberlake

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 - The Rise of Digital Reading Metals Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Metals Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metals Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metals Study Guide Answers
 - Personalized Recommendations
 - Metals Study Guide Answers User Reviews and Ratings
 - Metals Study Guide Answers and Bestseller Lists
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- Metals Study Guide Answers Enhanced eBook Features

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- Highlighting and Note-Taking Metals Study Guide Answers
- Interactive Elements Metals Study Guide Answers

8. Staying Engaged with Metals Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Metals Study Guide Answers

9. Balancing eBooks and Physical Books Metals Study Guide Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Metals Study Guide Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Metals Study Guide Answers

- Setting Reading Goals Metals Study Guide Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Metals Study Guide Answers

- Fact-Checking eBook Content of Metals Study Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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