

EDITABLE CONCERT TICKETS



Concert Tickets Guide

Frank Coffey



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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable
measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic
eg 1 HCl aq + NaOH aq → NaCl aq + H₂O l for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃ aq + KOH aq → KNO₃ aq + H₂O l for which ΔH = 57 kJ mol⁻¹

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react
together to produce one mole of water for example HCl aq + NaOH aq → NaCl aq + H₂O aq you can see that one mole of water is
produced when one mole of HCl or one mole of NaOH reacts

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