

ORGANIC CHEMISTRY

REACTIONS

FEATURES OF AN ORGANIC REACTION

Mechanism: Describes the overall reaction using a series of simple steps.

Stoichiometry: Calculate reactants and product masses using the balanced equation and molar masses.

Kinetics: The study of the reaction rate and mechanism.

Theoretical yield: Mass of product given by a complete reaction.

Yield: $\frac{\text{actual}}{\text{theoretical}}$ mass (theoretical yield).

Equilibrium: Reaction does not proceed to completion, instead, it reaches a balanced state of forward and reverse reactions.

Major reaction types:

- Acid
- Base
- Oxidation/Reduction
- Condensation
- Substitution (SN1, SN2)
- Ionic Reaction
- Elimination (E1, E2)
- Cationic
- Hydrolysis
- Addition
- Radical reaction

Important named reactions:

- **Halo-Alkane:** from alkyl halide
- **Friedel-Crafts:** alkyl or acyl group
- **Grignard:** alkyl or acyl group
- **Witt-Renbaum, Clemmensen:** reduce ketone to alkane
- **Wittig:** convert aldehyde/ketone to alkene

KINETICS AND REACTION MECHANISM



Transition state (TS): Maximum on the reaction coordinate curve; the least stable intermediate.

Activation Energy (E_a): Energy of the TS relative to the reactants. The change in enthalpy (ΔH) is $-E_a$ for exothermic, $+E_a$ for endothermic.

Hammett-Laurer postulate: The TS is more like the reactant or product that is closer in energy. endothermic TS is like the product, exothermic TS is like the reactant.

Kinetic vs. thermodynamic control: (a) and (b) describe Thermodynamic stability.

• If (a) is large and negative (exergonic), the product formation is likely controlled by "thermodynamic". Large K_{eq} corresponds to a large amount of product, relative to reactant.

• A large E_a may give rise to "kinetic" control; the energy of the TS controls the reaction, instead of the product/reactant thermodynamics.

Solvent effects: A solvent may stabilize an intermediate, decreasing E_a and increasing the rate of the reaction. Charge-complexes are stabilized by polar solvents.

ORGANIC ACID AND BASE

Acid:

- Electrophile: accepts δ^- ions acids
- Proton donor (Brønsted-Lowry acid)
- example: carboxylic acid

Base:

- Electrophile donor (Lewis base)
- Proton acceptor (Brønsted-Lowry base)
- example: amine

Factors enhancing acid strength (HA):

- Medium H-A bond
- Greater electronegativity of "A"
- Inductive effect of substituents on "A" (electron withdrawal enhances transfer)
- More "s" character in hybrid orbital involved in bond to group that is pulled
- Resonance stabilization conjugate base (A⁻)

Factors enhancing base strength:

- Resonance of conjugate acid/bases
- A base is a nucleophile. Electrons effects which shift electron density to the atom with the lone-pair increases base strength.

ALKANE

Properties:

- Hydrocarbon
- Weak intermolecular forces
- Non-polar, general formula C_nH_{2n+2}
- Saturated C-C-C (109°)

Nomenclature:

- Add "-ane" to prefix
- Lowest substituents by priority
- Halobutane: substituent halide for alk

Cycloalkane (C_nH_{2n}):

- Bicyclic: two fused or bridged rings
- $n = 2$: cyclopropane (highly strained)
- $n = 3$: cyclobutane (some flexibility)
- $n = 4$: cyclopentane (slight puckering)
- $n = 5$: cyclohexane: chair = stable conformation
- Axial: less stable, axial position, "perpendicular" to ring. Equatorial position, in ring "plane"
- Cis: two substituents on up position
- Trans: one up and one down



Synthesis:

- Hydrogenate alkene or alkyne (H_2 , Pt catalyst)
- Free-radical reaction of alkane
- Radical halobutane (Br, $h\nu$)
- Friedel-Crafts alkylation

Reactions:

- Combustion: alkane + $O_2 \rightarrow CO_2 + H_2O$
- Halogenation to halobutane ($Cl_2/h\nu$, light or heat)

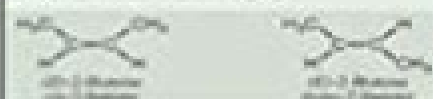
ALKENE $>C=C<$

Properties: Similar to alkane, non-polar, flammable

Nomenclature:

- Add "-ene" to prefix, Use E to denote C=C position
- Saturated C-C-C-C-C-C, unsaturated C=C-C-C
- Polyunsaturated fatty acid: 2 or more C=C
- Alkene: adjacent C-C=C
- Vinyl group: $H_2C=CH-$
- Methylene group: H_2C-
- Allyl group: $H_2C=CH-CH_2-$
- Vinyl halide: halide replaces H on C=C
- Conjugated alkenes: C=C and C=C (consecutive)
- Alkadiene, 2 conjugated C=C, example: butadiene, cis and trans isomers about C=C bonds
- Alkatriene, 3 conjugated C=C
- Aromatic: conjugated monocyclic compound, example: (H) benzene = benzene
- Aromatic: cyclic, non-cyclopentadiene anion, cyclopentadienyl anion (5 electrons)

Isomers: no free rotation of C=C



- E/Z: priority groups by atomic weight of - higher priority groups on the same side
- For monosubstituted: cis is less stable (more branching), the opposite, cis more stable
- **Markovnikov Rule:** from the less substituted alkene
- **Anti-Markovnikov addition:** H adds to C with more H's
- **Kation Electrophile:** from alkene with more substitution

Synthesis:

- Dehydration: alcohol (H^+ , heat) (elimination)
- Dehydrohalogenation: halobutane (heat, base)
- Dehalogenation: vic. dihalide (Zn, acetic acid)
- Hydrogenation: alkene
- sp^2 , Diels-Alder (H_2 , Pt catalyst)
- sp^2 , Diels-Alder (C=C, $h\nu$, $h\nu$)
- Wittig: aldehyde/ketone + phosphorus ylide

Reactions:

- Combustion: alkene
- Hydrogen to $2H^+$ alcohol (H^+ , H_2 , Pt) from alkene, see mechanism (Markovnikov)
- Hydrogen to alcohol: hydroboration-oxidation ($2H^+$, H_2 , H_2O_2 , OH^- , OH^- , OH^-) (anti-Markovnikov)
- Electrophilic: electrophile to alcohol
- Hydrohalogenation: alkene (Markovnikov)
- Halogenation: Br_2/H_2O vic. dihalobutane (Br_2/H_2O , anti addition)
- Halohydrin (Br_2 , H_2O , anti addition)
- Hydroboration: to form a 1,2-diol (alkene, OH^- , OH^- , anti addition)
- Oxidation: to carboxylic acid (KMnO₄, heat OH^-)
- Electrophile to ketone (Br_2 , Zn, H_2)
- Hydrogenation to alkane (H_2 , Pt, anti addition)
- Free-radical polymerization
- Alkylation: Reaction
- alkyl halogenation (Cl_2 , heat)
- Diels-Alder: cyclopentadiene from diene + alkene/diophile

Organic Chemistry Reactions Quickstudy Academic

Mark Jackson



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Organic Chemistry Reactions Mark Jackson,2001-09-19 This guide is packed with useful and up to date information regarding Organic Chemistry Reactions The laminated 4 page guide contains information on features of an organic reaction kinetics reaction mechanism organic acid base benzene arene alkyne alcohol and much more

ADVANCED ORGANIC CHEMISTRY: REACTIONS, MECHANISMS AND STRUCTURE, 4TH ED March,2006 Market_Desc Professors in Organic Chemistry Students in Organic Chemistry Organic Chemists Special Features The book Describes the structure of organic compounds including chemical bonding and stereochemistry Reviews general reaction mechanisms including ordinary reactions and photochemical reactions Includes a survey of reactions arranged by reaction type and by which bonds are broken and formed Includes IUPAC s newest system for designating reaction mechanisms Features an index to the methods used for preparing given types of compounds Contains more than 15 000 references 5 000 new to this edition to original papers About The Book The book covers the three fundamental aspects of the study of organic chemistry reactions mechanisms and structure Part One explores the structure of organic compounds providing the necessary background for understanding mechanisms Part Two discusses reactions and mechanisms Organized by reaction type each of these chapters discusses the basic mechanisms along with reactivity and orientation as well as the scope and mechanisms of each reaction

Understanding Organic Reaction Mechanisms Adam Jacobs,1997-07-17 This book describes the principles that govern chemical reactivity and shows how these principles can be used to make predictions about the mechanisms and outcomes of chemical reactions Molecular orbital theory is used to provide up to date explanations of chemical reactivity in an entirely nonmathematical approach suited to organic chemists A valuable section explains the use of curly arrows vital for describing reaction mechanisms An entire chapter is devoted to exploring the thought processes involved in predicting the mechanisms of unfamiliar reactions Each chapter is followed by a summary of the important points and a selection of problems to help the reader make sure that the material in that chapter has been assimilated The book concludes with a comprehensive glossary of technical terms This text will be of interest to first and second year chemistry undergraduates studying organic chemistry

Organic Chemistry Fundamentals BarCharts, Inc,2015

A Self-study Guide to the Principles of Organic Chemistry Jiben Roy,2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future

study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix *Advanced Organic Chemistry* Jerry March,1985 This survey of advanced chemistry covers virtually all the useful reactions 600 all told with the scope limitations and mechanism of each described in detail Extensive general sections on the mechanisms of the important reaction types and five chapters on the structure and stereochemistry of organic compounds and reactive intermediates are included as well Of the more than 10 000 references included 5 000 are new in this edition *Organic Chemistry Reactions* Speedy Publishing,2014-09-03 Students of organic chemistry are expected to consume much information in a relatively short period of time Most have had no clue to the expanse of knowledge that organic chemistry explores Students are required to memorize elements and molecules that are commonly used in organic chemistry Additionally they are required to memorize formulas and chemical reactions which is clearly the most difficult part of the course Having an organic chemistry reaction study guide can help the student by supplying a quick reference to the most commonly used reactions The guide can be reviewed when the student has some down time *Organic Chemistry Reactions: A Study Guide* Cybellium,2024-10-26 Designed for professionals students and enthusiasts alike our comprehensive books empower you to stay ahead in a rapidly evolving digital world Expert Insights Our books provide deep actionable insights that bridge the gap between theory and practical application Up to Date Content Stay current with the latest advancements trends and best practices in IT AI Cybersecurity Business Economics and Science Each guide is regularly updated to reflect the newest developments and challenges Comprehensive Coverage Whether you re a beginner or an advanced learner Cybellium books cover a wide range of topics from foundational principles to specialized knowledge tailored to your level of expertise Become part of a global network of learners and professionals who trust Cybellium to guide their educational journey www.cybellium.com

Advanced Organic Chemistry: Reactions And Mechanisms Maya Shankar Singh,Singh,2004-09 Advanced Organic Chemistry Reactions and Mechanisms covers the four types of reactions substitution addition elimination and rearrangement the three types of reagents nucleophiles electrophiles and radicals and the two effects electroni *Advanced Organic Chemistry* Francis A. Carey,1990 *Advanced Organic Chemistry* Francis A. Carey,1977 **Organic Chemistry Reactions (Speedy Study Guide)** Speedy Publishing LLC,2014-09-04 Students of organic chemistry are expected to consume much information in a relatively short period of time Most have had no clue to the expanse of knowledge that organic chemistry explores Students are required to memorize elements and molecules that are commonly used in organic

chemistry Additionally they are required to memorize formulas and chemical reactions which is clearly the most difficult part of the course Having an organic chemistry reaction study guide can help the student by supplying a quick reference to the most commonly used reactions The guide can be reviewed when the student has some down time

Named Organic Reactions Thomas Laue, Andreas Plagens, 2005-08-19 This Second edition contains concise information on 134 carefully chosen named organic reactions the standard set of undergraduate and graduate synthetic organic chemistry courses Each reaction is detailed with clearly drawn mechanisms references from the primary literature and well written accounts covering the mechanistic aspects of the reactions and the details of side reactions and substrate limitations For the 2nd edition the complete text has been revised and updated and four new reactions have been added Baylis Hillmann Reaction Sonogashira Reaction Pummerer Reaction and the Swern Oxidation and Cyclopropanation An essential text for students preparing for exams in organic chemistry

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web aug 26 2011 in rational recovery you learn to separate your thinking higher rational self from your addictive voice or what jack trimpey metaphorically calls the beast the beast is your animal midbrain it is the part of your brain that wants only immediate pleasure drinking drugging eating screwing with absolutely no concern for any

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