

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



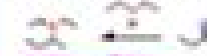
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Ions



Bond and Reaction Rates



Bonding in Molecules



There are two types of bonding

Definition of ionic bond



Definition of covalent bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.

- **Single bond** is a sigma bond.
- **Double bond** consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity:

- **Nonpolar covalent bonds** are formed between two atoms with similar electronegativity.
- **Polar covalent bonds** are formed between two atoms with different electronegativity.



- If the difference in electronegativity is less than 0.5, the bond is nonpolar.
- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.
- If the difference in electronegativity is greater than 1.7, the bond is ionic.

- A difference of 1.7 or higher indicates that the electrons are not shared equally and a ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the individual bonds are polar. A good example is carbon dioxide.



Consider the vector sum of the dipole moments for carbon dioxide.

Bond Length and Bond Strength

The longer the bond, the weaker the bond.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

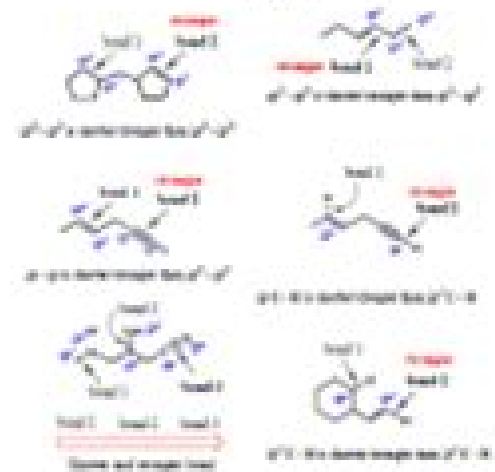


The longer the bond, the weaker the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	410 kJ/mol	602 kJ/mol	837 kJ/mol

Higher s character means shorter and stronger bond

Examples



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N. Sue Barnes



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