

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



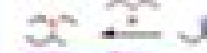
- **Covalent bonds** are formed by sharing of electrons between two atoms.



### The Relative Strength of Bonds



### Bond and Reaction Rates



### Bonding in Molecules



### There are two types of bonding

Definition of ionic bond



## Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.

- **pi bonds** are formed by side-on overlap of orbitals.

- **One single bond is a sigma bond.**

- **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond). For example, the C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond). For example, the C-O and C-N bonds.

- A difference of 1.7 or higher is ionic (the electrons are not shared).

## Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ<sup>-</sup>) and the less electronegative atom has a partial positive charge (δ<sup>+</sup>).

- **δ<sup>-</sup> (delta minus)** indicates an electron-deficient atom.

- **δ<sup>+</sup> (delta plus)** indicates an electron-rich atom.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

## Bond Length and Bond Strength

The length of a bond is important.

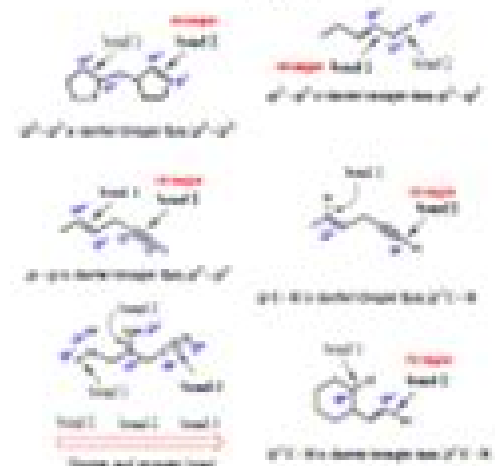


The length of a bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp <sup>3</sup> -C-C	sp <sup>2</sup> -C-C	sp-C-C
Hybridization	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	160 kcal/mol
Hybridization	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
C-H Bond	sp <sup>3</sup> -H	sp <sup>2</sup> -H	sp-H
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

## Examples



# Organic Chem Final Study Guide

**R Barnett**



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