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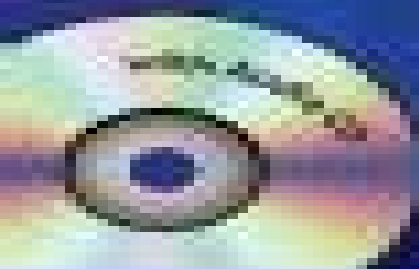
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# Objective Proficiency

Workbook  
with answers

For revised exams from March 2013



Robert Sunderland    Erica Miskerton    Second Edition

# Objective Proficiency Workbook With Answers With Audio Cd

**Daniel F McAuley**



## **Objective Proficiency Workbook With Answers With Audio Cd:**

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web determination of an unknown quantity of aspirin using back titration back ground theory many reactions are slow or present unfavourable equilibria for direct titration aspirin is a weak acid therefore it can react with a base such as sodium hydroxide  $\text{NaOH}$  this is an example of an acid base titration see reaction 1

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web why is back titration used for aspirin consider using titration to measure the amount of aspirin in a solution using back titration the end point is more easily recognised in this reaction as it is a reaction between a strong base and a strong acid

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web determination of aspirin using back titration this experiment is designed to illustrate techniques used in a typical indirect or back titration you will use the  $\text{NaOH}$  you standardized last week to back titrate an aspirin solution and determine the concentration of aspirin in a typical analgesic tablet you will be graded on your accuracy

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web topic determination of aspirin in tablets using back titration aim to calculate and compare the active pharmaceutical ingredient  $\text{API}$  in different commercially available aspirin tablets of the same batch using titration technique standard  $\text{s}$  indicator  $\text{s}$  5 1 8 b1

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web what is back titration it is basically an analytical technique in chemistry which is performed backwards in the method that is a user needs to find the concentration of a reactant of a given unknown concentration by reacting it with an excess volume of another reactant of a known concentration

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web second you will titrate a sample of your aspirin acetylsalicylic acid with the standardized  $\text{NaOH}$  to determine the moles of acid in a given weight of your product this will allow you to assess its purity assuming the aspirin is not contaminated with other acids the titration allows you to quantitatively determine the purity of your aspirin



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web 4 6 back titration 1 the active ingredient in aspirin tablets is acetylsalicylic acid  $C_9H_8O_4$  the acetylsalicylic acid content of an aspirin tablet can be determined using a back titration four aspirin tablets were crushed and added to 25.0 cm<sup>3</sup> of 1.00 mol l<sup>-1</sup> sodium hydroxide solution

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web pubchemdatabase aspirin cid 2244 here back titration was done to obtain the concentration of aspirin from known aspirin tablets amount of product expected was 0.5g measured yield obtained was 0.335g which shows there was an error in the determination

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web of this study is to analysis of aspirin and compares both direct titration and back titration techniques to the standard hplc technique in order to find alternative easy and cheaper procedure to assess the quality of active components of medications aspirin in

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