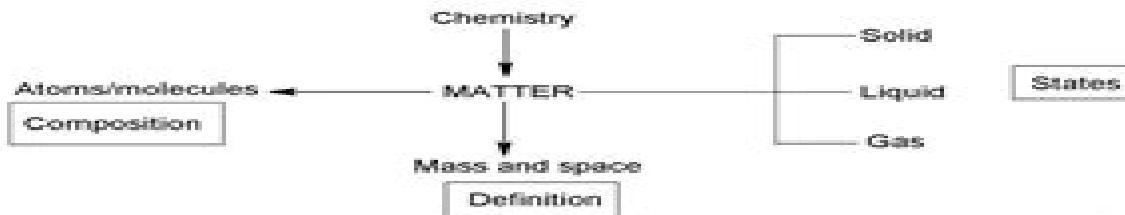


O level Chemistry Notes



Atoms: Building block of matter. It may or may not exist free in nature
Examples: H, He, Na, C

Shells = electrons
 Mass = 0
 Charge = -1



Nucleus = Protons + Neutrons
 Charge = +1
 Mass = 1

Nucleus: It is central part of an atom.

Shells: Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell = $2n^2$ Where n = no of shell.

1st shell = 2es, 2nd shell = 8es, 3rd shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

Atomic number:

Number of protons present in the nucleus of an atom of an element.

Examples: 11Na, 6C, 8O

Atomic mass or Mass number or nucleon number: It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples: ²³Na, ¹²C, ¹⁶O

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Frances Talaska Fischbach

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Table of Contents O Level Chemistry Sample Chapter 1

1. Understanding the eBook O Level Chemistry Sample Chapter 1
 - The Rise of Digital Reading O Level Chemistry Sample Chapter 1
 - Advantages of eBooks Over Traditional Books
2. Identifying O Level Chemistry Sample Chapter 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an O Level Chemistry Sample Chapter 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from O Level Chemistry Sample Chapter 1
 - Personalized Recommendations
 - O Level Chemistry Sample Chapter 1 User Reviews and Ratings
 - O Level Chemistry Sample Chapter 1 and Bestseller Lists
5. Accessing O Level Chemistry Sample Chapter 1 Free and Paid eBooks
 - O Level Chemistry Sample Chapter 1 Public Domain eBooks
 - O Level Chemistry Sample Chapter 1 eBook Subscription Services
 - O Level Chemistry Sample Chapter 1 Budget-Friendly Options
6. Navigating O Level Chemistry Sample Chapter 1 eBook Formats

- ePub, PDF, MOBI, and More
- O Level Chemistry Sample Chapter 1 Compatibility with Devices
- O Level Chemistry Sample Chapter 1 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of O Level Chemistry Sample Chapter 1
- Highlighting and Note-Taking O Level Chemistry Sample Chapter 1
- Interactive Elements O Level Chemistry Sample Chapter 1

8. Staying Engaged with O Level Chemistry Sample Chapter 1

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers O Level Chemistry Sample Chapter 1

9. Balancing eBooks and Physical Books O Level Chemistry Sample Chapter 1

- Benefits of a Digital Library
- Creating a Diverse Reading Collection O Level Chemistry Sample Chapter 1

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine O Level Chemistry Sample Chapter 1

- Setting Reading Goals O Level Chemistry Sample Chapter 1
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of O Level Chemistry Sample Chapter 1

- Fact-Checking eBook Content of O Level Chemistry Sample Chapter 1
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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