

# NCERT Solutions for Class-XI Chemistry

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## Chapter-2

### NCERT Chemistry Class 11

1. Calculate the molecular mass of the following:
  - (i) Calculate the number of electrons which will together weigh one gram.
  - (ii) Calculate the mass and charge of one mole of electrons.
1. (i) Mass of one electron =  $9.10939 \times 10^{-31}$  kg  
 $\therefore$  Number of electrons that weigh  $9.10939 \times 10^{-31}$  kg = 1  
 $\therefore$  Number of electrons that will weigh 1 g ( $1 \times 10^{-3}$  kg)  
$$= \frac{1}{9.10939 \times 10^{-31} \text{ kg}} \times (1 \times 10^{-3} \text{ kg})$$
$$= 0.1098 \times 10^{-3} + 31$$
$$= 0.1098 \times 10^{28}$$
$$= 1.098 \times 10^{27}$$
  - (ii) Mass of one electron =  $9.10939 \times 10^{-31}$  kg  
Mass of one mole of electron =  $(6.022 \times 10^{23}) \times (9.10939 \times 10^{-31} \text{ kg})$   
 $= 5.48 \times 10^{-7} \text{ kg}$   
Charge on one electron =  $1.6022 \times 10^{-19}$  coulomb  
Charge on one mole of electron =  $(1.6022 \times 10^{-19} \text{ C}) (6.022 \times 10^{23})$   
 $= 9.65 \times 10^4 \text{ C}$
2.
  - (i) Calculate the total number of electrons present in one mole of methane.
  - (ii) Find (a) the total number and (b) the total mass of neutrons in 7 mg of  $^{14}\text{C}$ . (Assume that mass of a neutron =  $1.675 \times 10^{-27}$  kg).
  - (iii) Find (a) the total number and (b) the total mass of protons in 34 mg of  $\text{NH}_3$  at STP. Will the answer change if the temperature and pressure are changed?
2. No the answer will not vary with the change in temperature and pressure because the number of subatomic particles like protons, neutrons and electrons is fixed for each and every element and it does not vary with temperature and pressure.
3. How many neutrons and protons are there in the following nuclei?  
 $^{12}_6\text{C}$ ,  $^{16}_8\text{O}$ ,  $^{24}_{12}\text{Mg}$ ,  $^{56}_{26}\text{Fe}$ ,  $^{88}_{38}\text{Sr}$
3.  $^{13}_6\text{C}$ : Atomic mass = 13  
Atomic number = Number of protons = 6  
Number of neutrons = (Atomic mass) – (Atomic number)  
 $= 13 - 6 = 7$   
 $^{16}_8\text{O}$ :  
Atomic mass = 16  
Atomic number = 8  
Number of protons = 8

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**Priti Singhal**



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