

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Heat Engines**: Devices that convert heat into work. The efficiency of a heat engine is given by:

$$\eta = \frac{W}{Q_{in}}$$

 where η is the efficiency, W is the work done, and Q_{in} is the heat input.
 - **Carnot Cycle**: The most efficient cycle possible between two isotherms. The efficiency is given by:

$$\eta_{Carnot} = 1 - \frac{T_c}{T_h}$$

 where T_c is the temperature of the cold reservoir and T_h is the temperature of the hot reservoir.



- **Refrigerators and Heat Pumps**: Devices that transfer heat from a cold reservoir to a hot reservoir. The coefficient of performance (COP) is given by:

$$COP = \frac{Q_c}{W}$$

 where Q_c is the heat removed from the cold reservoir and W is the work done.
 - **Entropy**: A measure of the disorder or randomness of a system. The change in entropy for a reversible process is given by:

$$\Delta S = \frac{Q_{rev}}{T}$$

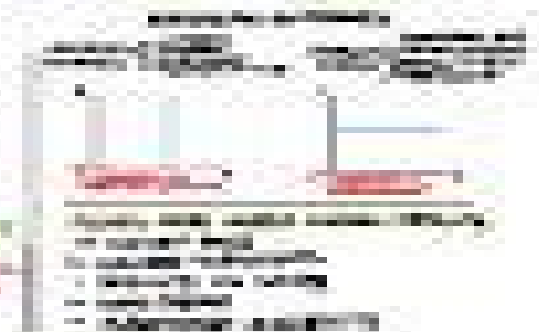
 where Q_{rev} is the heat added reversibly and T is the temperature.



- **Statistical Mechanics**: The study of the behavior of large numbers of particles. The Boltzmann constant (k_B) relates the average kinetic energy of particles to the temperature:

$$\frac{1}{2} m \langle v^2 \rangle = \frac{3}{2} k_B T$$

 where m is the mass of a particle, $\langle v^2 \rangle$ is the mean square velocity, and T is the temperature.



- **Phase Transitions**: Changes in the state of matter (solid, liquid, gas, plasma). The phase diagram shows the regions of stability for each phase as a function of temperature and pressure.



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