

# 

- **Thermodynamics** is the study of energy and its transformations.  
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.  

$$\Delta U = Q - W$$
  
 where  $\Delta U$  is the change in internal energy,  $Q$  is the heat added to the system, and  $W$  is the work done by the system.  
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.  

$$\Delta S \geq 0$$
  
 where  $\Delta S$  is the change in entropy.



- **Enthalpy** ( $H$ ) is a thermodynamic property that represents the total energy of a system, including internal energy and the energy required to create the system's volume against the external pressure.  

$$H = U + PV$$
  
 where  $U$  is internal energy,  $P$  is pressure, and  $V$  is volume.  
 - **Enthalpy Change** ( $\Delta H$ ) is the change in enthalpy during a process.  

$$\Delta H = H_{\text{final}} - H_{\text{initial}}$$
  
 - **Exothermic Process**: A process that releases heat to the surroundings, resulting in a negative  $\Delta H$ .  
 - **Endothermic Process**: A process that absorbs heat from the surroundings, resulting in a positive  $\Delta H$ .



- **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.  

$$S = 0 \text{ at } T = 0 \text{ K}$$
  
 - **Gibbs Free Energy** ( $G$ ) is a thermodynamic potential that measures the maximum reversible work that can be performed by a system at constant temperature and pressure.  

$$G = H - TS$$
  
 where  $H$  is enthalpy,  $T$  is temperature, and  $S$  is entropy.  
 - **Gibbs Free Energy Change** ( $\Delta G$ ) is the change in Gibbs free energy during a process.  

$$\Delta G = \Delta H - T\Delta S$$
  
 - **Spontaneous Process**: A process that occurs without the need for external energy input, resulting in a negative  $\Delta G$ .  
 - **Non-spontaneous Process**: A process that requires external energy input to occur, resulting in a positive  $\Delta G$ .



# Microeconomics Study Guide Final

**McConnell, William B. Walstad**



## Microeconomics Study Guide Final:

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