

STUDY GUIDE - ATOMIC STRUCTURE

The current atomic theory has come about over the past 2000 years. In the beginning, Democritus proposed the idea that all things were made of particles too small to see. He was laughed at.

In the 1800's John Dalton proposed the idea of the "Atomic Theory". He had 5 theories, 3 of which are still believed to-day. They are:

1. All matter is composed of extremely small particles too small to see.
2. In reactions, atoms separate, combine or rearrange.

3. Different atoms can combine in simple, whole number ratios (to form compounds).

Dalton believed that atoms could not be broken into smaller particles. That belief was proven false in the 1890's when JJ Thomson discovered the electron. His idea was that the atom had a negative charge, but he thought the electrons were just found within an atom, like plum pudding, not circling the atom. He made his discovery using a cathode ray tube.

In 1911 Ernest Rutherford shot alpha particles through a piece of gold foil. He found that some of the particles went through the gold foil, but most bounced off. He concluded that atoms had a densely packed, positively charged center - called the nucleus.

In 1913 Neils Bohr found that electrons are not actually in the nucleus, but circle it instead. He proposed the idea that electrons were found in energy levels. He believed the electrons circled the nucleus the way the planets orbit the sun. His theory has been revised since then - now electrons are believed to be found in orbitals around the nucleus, called the quantum theory.

Atomic Structure

Terms:

1. **Atomic number:** # of protons (whole number on the PT)
2. **Average Atomic Mass:** mass according to the periodic table (decimal number on the PT)
3. **Mass number:** determined by adding the "#p" to the "# n".
4. **Ions:** When an atom has a charge - comes from unequal numbers of protons and electrons. This occurs when electrons are gained or lost in a reaction.
5. **Isotope:** atoms of the same element have different atomic masses due to different numbers of neutrons.
6. **Nuclear symbol:** Show the element's symbol, the atomic mass & the atomic number. Mass goes on the top left of the symbol, number goes on the bottom left of the symbol. $^{197}_{\text{Ar}}$
7. **Bohr Model:** Shows number of electrons in each energy level around the nucleus.

Answer the following questions regarding the discovery of the atom & its parts.

- A. Identify 3 conclusions of the gold foil experiment.
- B. Which experiment discovered the electron?
- C. Why was the electron discovered first?
- D. Which parts of the atomic theory are still considered true to-day?

Know the parts of the atom:

- A. Which are in the nucleus?
- B. Which have a mass of 1 u?
- C. Which does not contribute to the mass of an atom?
- D. What is the charge on each subatomic particle?
- E. What is the job of each subatomic particle?
- F. If an atom loses 3 electrons, what is its charge and name?
- G. How are isotopes of the same element different?
- H. If an atom gains an electron in a reaction, what are its charge and its name?

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Austin Dulebohn

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 - The Rise of Digital Reading Nucleus Study Guide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Nucleus Study Guide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nucleus Study Guide Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nucleus Study Guide Answer
 - Personalized Recommendations
 - Nucleus Study Guide Answer User Reviews and Ratings
 - Nucleus Study Guide Answer and Bestseller Lists
5. Accessing Nucleus Study Guide Answer Free and Paid eBooks
 - Nucleus Study Guide Answer Public Domain eBooks
 - Nucleus Study Guide Answer eBook Subscription Services
 - Nucleus Study Guide Answer Budget-Friendly Options

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- ePUB, PDF, MOBI, and More
- Nucleus Study Guide Answer Compatibility with Devices
- Nucleus Study Guide Answer Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Nucleus Study Guide Answer
- Highlighting and Note-Taking Nucleus Study Guide Answer
- Interactive Elements Nucleus Study Guide Answer

8. Staying Engaged with Nucleus Study Guide Answer

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Nucleus Study Guide Answer

9. Balancing eBooks and Physical Books Nucleus Study Guide Answer

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Nucleus Study Guide Answer

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Nucleus Study Guide Answer

- Setting Reading Goals Nucleus Study Guide Answer
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Nucleus Study Guide Answer

- Fact-Checking eBook Content of Nucleus Study Guide Answer
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements

- Interactive and Gamified eBooks

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