

# Chemical Bonding

## 1. OCTET RULE (LEWIS-KOSSEL RULE)

Noble gases have 8 electrons in their outermost shell (complete octet) and outermost configuration is  $ns^2p^6$ . Every atom has a tendency to complete its octet by losing or gaining or by sharing electrons.

### 1.1 Exceptions of Octet Rule

(i) **Electron-Deficient Compounds** : Compounds having less than 8 electrons in the outermost shell of the central atom. e.g.  $\text{BF}_3$ ,  $\text{BeCl}_2$ ,  $\text{AlCl}_3$ .

(ii) **Expansion of Octet** : The compounds having more than 8 electrons in the outermost shell of the central atom. e.g  $\text{PCl}_5$ ,  $\text{SF}_6$ .

(iii) **Transition Metal Ions** : Transition metal ions have 9 to 18 electrons in their outermost shell. e.g.  $\text{Cr}^{+3}$ ,  $\text{Mn}^{+2}$ ,  $\text{Cu}^{+}$ .

(iv) H and Li atoms attain configuration like He by gaining and losing one electron, respectively. e.g.  $\text{H}^{+1}$ ,  $\text{Li}^{+1}$ .

(v) **Odd Electron Bond** : Compounds having unpaired electron are also exceptions to the octet rule. The molecules having odd electrons are  $\text{NO}_2$ ,  $\text{NO}_3^-$ ,  $\text{ClO}_4^-$ ,  $\text{O}_2^{+1}$  etc.

**Ex.1** Octet rule is not followed in:

$$\text{[1]} \text{ Ti}^{+3} \quad \text{[2]} \text{ Na}^{+} \quad \text{[3]} \text{ Ca}^{+2} \quad \text{[4]} \text{ K}^{+1} \quad \text{Ans. [1]}$$

**Ques.**  $Tl^{+3}$  has 9 electron in its outermost shell.

**Ex.2** Which of the following compounds has the central atom with incomplete octet ?

[1]  $\text{NH}_3$  [2]  $\text{BCl}_3$  [3]  $\text{CCl}_4$  [4]  $\text{PCl}_5$  Ans. [2]

**Sol.** B has 6 electron in  $\text{BCl}_3$ .

**Ex.3** An element forms cation by losing electron. The configuration of the cation may be like that of :

[1] inert gas      [2] pseudo inert gas      [3] inert pair      [4] all above      Ans. [4]

**Sol.** Inert gas — 8 electron in outermost shell. Pseudo inert gas — 18 electron in outermost shell.

## 1.2 Inert Pair Configuration

The heavy elements of IIIA, IVA, and VA groups form cations having charge equal to group number and two less than group number, e.g.  $Ti^{+3}$ ,  $In^{+3}$ ,  $Sn^{+2}$ ,  $Pb^{+2}$ . In these elements due to more penetration power of s orbitals, the atom generally does not lose  $e^-$  present in  $ns$  orbital, so only  $np$  electrons are lost by the atom. This effect is called "inert pair effect".

## 2. IONIC BOND

(i) Atoms get established by ion formation in formation of an ionic bond. One atom forms a cation by losing electrons and other forms anion by accepting electrons.

(ii) The octets of the atoms are completed by transfer of electrons.

(iii) Cations and anions are bonded by electrostatic force of attraction in ionic bond.

(iv) Ionic bond is formed between metals (cations) and nonmetals (anions).

(v) During the formation of an ionic bond, cation can attain one of the following configurations.

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Datta,2019-02-25 The concept of a chemical bond evolved from a variety of experimental observations It became useful to understand at times even predict the molecular structure reactivity and mechanism of chemical reactions Every aspect of the concept of bonding received a quantitative interpretation from the advent of quantum mechanics and its application to chemistry In Lectures on Chemical Bonding and Quantum Chemistry the reader will find a comprehensive discourse on the basic interpretation of the chemical bond as well as current understanding in terms of a dancing molecule that not only travels rotates and pulsates around an equilibrium molecular structure but also interacts and collides with other molecules thereby transferring linear and angular momentum characteristics and adjusting total energies One will also find a thorough survey of quantum mechanical methodologies for calculation of molecular characteristics in specific states and their changes under spectroscopic transitions tunneling electron and proton transfer phenomena and so on Guides to more advanced levels of theory are also provided Atomic Structure and Chemical Bonding, a Non-mathematical Introduction Fritz Seel,1963

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