

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bonding Interactions



Bonding Interactions



There are two types of bonding interactions



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two double bonds** consist of one sigma and one pi bond.



• **pi bonds** are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **If the difference in electronegativity is less than 0.5**, the electrons are shared equally (nonpolar).

• **If the difference in electronegativity is between 0.5 and 1.7**, the electrons are shared unequally (polar).

• **A difference of 1.7 or higher** is so high that the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



It is a vector quantity.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

• **δ⁻ (delta minus)** indicates an electron deficient atom.

• **δ⁺ (delta plus)** indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar.



• **Nonpolar molecules** have no net dipole moment.

Bond Length and Bond Strength

• **Single bonds** are the longest.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

• **Double bonds** are shorter than single bonds.

Bond	C=C	C≡C
Length	134 pm	120 pm
Strength	614 kJ/mol	839 kJ/mol

• **Triple bonds** are the shortest.

	sp ³ C	sp ² C	sp C
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	120 pm
Strength	338 kJ/mol	614 kJ/mol	839 kJ/mol
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	120 pm
Strength	338 kJ/mol	614 kJ/mol	839 kJ/mol

• **Triple bonds** are the shortest.

Examples



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