

# Study Guide A

## Answer Key

### SECTION 1. CHEMICAL ENERGY AND ATP

1. molecule; food molecules
2. high-energy; lower-energy
3. phosphate group
4. a; d; b; c
5. b; c
6. c; d
7. a; f
8. chemical energy; light energy
9. three; two
10. chemical; molecules

### SECTION 2. OVERVIEW OF PHOTOSYNTHESIS

1. chemical energy
2. light energy; chemical energy
3. visible light
4. photosynthesis
5. stroma; grana
6. grana
7. Equation:  $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$
8.  $6\text{CO}_2$  and  $6\text{H}_2\text{O}$
9.  $6\text{O}_2$  and  $\text{C}_6\text{H}_{12}\text{O}_6$
10. b
11. require light; energy
12. do not require light; sugars
13. Refer to Figure 2.2 for a visual answer.  
The four steps are: d, a, b, c
14. energy; sugars
15. green
16. independent

### SECTION 3. PHOTOSYNTHESIS IN DETAIL

1. capture; transfer
2. molecules
3. ATP; NADPH
4. b, g, d, f, c, e, a
5. carbon dioxide
6. 1 b; 2 d; A a; 3 c; 4 c
7. proteins
8. synthesizes
9. with no beginning or end.

### SECTION 4. OVERVIEW OF CELLULAR RESPIRATION

1. energy; oxygen
2. oxygen
3. mitochondria
4. glucose; ATP
5. b; c
6. a; c
7. Equation:  $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$
8.  $6\text{O}_2$  and  $\text{C}_6\text{H}_{12}\text{O}_6$
9.  $6\text{CO}_2$  and  $6\text{H}_2\text{O}$
10. d
11. Refer to Figure 4.2 for a visual answer.  
The four steps are: d, b, a, c
12. sugar
13. oxygen
14. break down

### SECTION 5. CELLULAR RESPIRATION IN DETAIL

1. glucose
2. Pyruvate and NADH; ATP
3. two; four
4. c
5. c, b, a, d, e
6. inner
7. electrons; hydrogen ions
8. c, d, b, a
9. electrons

### SECTION 6. FERMENTATION

1. ATP
2. electrons
3. hard exercise
4. oxygen
5. two
6. For a visual answer, refer to the lactic acid fermentation diagram in Section 6. The order of the four steps is: d, b, c, a
7. For a visual answer, refer to the alcoholic fermentation diagram in Section 6. The order of the four steps is: c, a, d, b

# Monster Answers Key Study Guide

**Christian Drosten**



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web question 2 what are the oxidation number of the underlined elements in each of the following and how do you rationalise your results answer a in  $\text{KIO}_3$  since the oxidation number of K is 1 therefore the average oxidation number of iodine is  $\frac{1}{3}$  but the oxidation number cannot be fractional

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