

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding between atoms in a molecule.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



Regions (p) and pi (π) bonds

- Covalent bonds are classified as sigma (σ) and pi (π).
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- Very simple bond is a covalent bond.
- Non-overlapping bond consists of two regions and two pi (π) bonds.



pi bonds are stronger than sigma.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a C-C and C-H bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared in a region of N, O, F and other halogens.
- A difference of 1.7 or higher one of both the atoms can have a larger share and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles

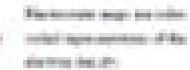
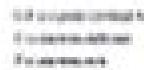


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Electronegativity is inversely proportional to the dipole moment.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a distance and area.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of molecular dipole moments. The dipole moment is measured. A good example is water, dipole.



Water has a net molecular dipole of 1.85 D.

Average values of C-C and C-H bond depending on electronegativity

| C-C Bond | $\text{H}_2\text{C}-\text{CH}_3$ | $\text{H}_2\text{C}=\text{CH}_2$ | $\text{CH}_3\text{CH}_2\text{CH}_3$ |
|-------------------|----------------------------------|----------------------------------|-------------------------------------|
| Electronegativity | 1.37 D | 1.37 D | 1.37 D |
| Length | 1.57 Å | 1.35 Å | 1.53 Å |
| Strength | 0.16 kcal/mol | 0.17 kcal/mol | 0.16 kcal/mol |
| Hybridization | sp^3 | sp^2 | sp^3 |
| Difference | 0.57 D | 0.57 D | 0.57 D |
| Length | 1.10 Å | 1.09 Å | 1.10 Å |
| Strength | 0.16 kcal/mol | 0.17 kcal/mol | 0.16 kcal/mol |

Dipole + dipole + dipole = strong bond

Example



Water + water = $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$

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Bond Length and Bond Strength

Reactions of bond length and strength



Bond: $\text{H}-\text{H}$ $\text{H}-\text{O}$ $\text{H}-\text{S}$ $\text{H}-\text{Cl}$

Length: 1.10 Å 1.27 Å 1.47 Å 1.70 Å

Strength: 0.16 kcal/mol 0.17 kcal/mol 0.16 kcal/mol 0.16 kcal/mol

Hydrogen bond: $\text{H}-\text{O}$ $\text{H}-\text{S}$ $\text{H}-\text{Cl}$

Reactions of bond length and strength

Hydrogen bond: $\text{H}-\text{O}$ $\text{H}-\text{S}$ $\text{H}-\text{Cl}$

Hydrogen bond: $\text{H}-\text{O}$ $\text{H}-\text{S}$ $\text{H}-\text{Cl}$

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