

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: $\text{NH}_4^+ + \text{H}_3\text{N}^- \rightarrow \text{NH}_3 + \text{NH}_4^+$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metal bond



The difference between ionic and covalent bonds



Covalent bond



Hydrogen bond



Metal bond



Electrostatic interaction



Electrostatic interaction



Electrostatic interaction



Electrostatic interaction



A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles



Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

1. 0.000000 N
2. 0.000000 N
3. 0.000000 N



Electrostatic map of molecular dipole moment of the water molecule.

- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The water molecule has a partial negative charge (δ-) and the oxygen atom has a partial positive charge (δ+).
- The dipole moment is the difference between the two.
- The dipole moment is shown as a vector with arrows.

Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. The dipole moment of the molecule is measured. A good example is water.



Water has a net molecular dipole moment of 1.85 D.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of atoms, polar bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. This is a nonpolar bond. (For example, the C-C bond in ethane.)
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. (Nonpolar bonds in our atoms have carbon as a ligand in N, O, F and other halogens.)
- A difference of 1.7 or higher one of the atoms has a stronger bond and a nonpolar bond is formed.

Bond Length and Bond Strength

Reactions of bond length and strength



Bond: H-H H-O-H H-Cl H-F
Length: 1.07 Å 1.21 Å 1.47 Å 1.66 Å

Strength: 100 kJ/mol 100 kJ/mol 100 kJ/mol 100 kJ/mol

Water is a polar molecule due to its smaller and stronger bond.

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Reactions of bond length and strength

Average values of C-C and C-H bond depending on the molecule

C-C bond: $1.54\text{--}1.56\text{ \AA}$ $1.52\text{--}1.54\text{ \AA}$ $1.50\text{--}1.52\text{ \AA}$

Hydrogen bond: $1.15\text{--}1.17\text{ \AA}$ $1.13\text{--}1.15\text{ \AA}$ $1.11\text{--}1.13\text{ \AA}$

Length: 1.17 \AA 1.15 \AA 1.13 \AA

Strength: 100 kJ/mol 100 kJ/mol 100 kJ/mol

Hydrogen bond: 1.15 \AA 1.13 \AA $1.11\text{--}1.13\text{ \AA}$

Strength: $10\text{--}15\text{ kJ/mol}$ $10\text{--}15\text{ kJ/mol}$ $10\text{--}15\text{ kJ/mol}$

Length: 1.13 \AA 1.11 \AA 1.09 \AA

Strength: 100 kJ/mol 100 kJ/mol 100 kJ/mol

Water is a polar molecule due to its smaller and stronger bond

Example



Water is a polar molecule due to its smaller and stronger bond.



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Water is a polar molecule due to its smaller and stronger bond.



Water is a polar molecule due to its smaller and stronger bond.



Water is a polar molecule due to its smaller and stronger bond.

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