

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of halogen bond: covalent and ionic.

- Covalent bonds are formed by the sharing of electrons from one orbital of one atom (nucleophile) and one orbital of another atom (electrophile).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



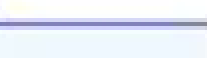
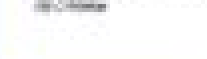
Bond and Electron Sharing



Bonding in Molecules



There are two types of halogen bond: covalent and ionic.



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

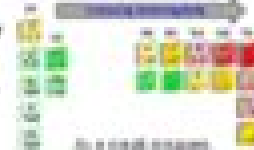
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 0.5 or higher can be high for the electrons are not shared equally (polar bond).

- A difference of 1.7 or higher can be high for the electrons are not shared equally (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- δ⁻ (delta minus) indicates an electron-deficient atom.
- δ⁺ (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

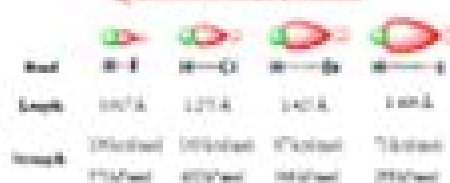
The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The length of a bond is important.

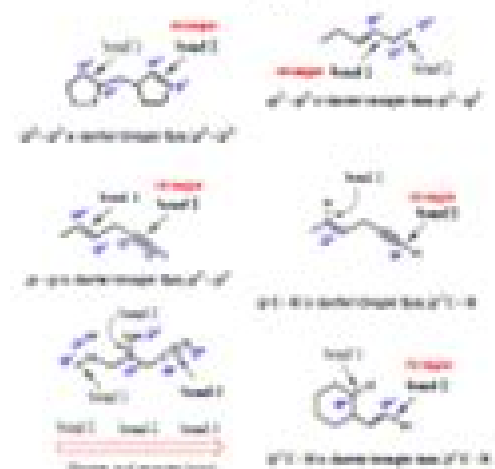


The length of a bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	472 kJ/mol	505 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	413 kJ/mol	439 kJ/mol	505 kJ/mol

Higher s character means stronger bond

Examples



Organic Chemistry I Study Guide

Morrison



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