

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding between atoms in a molecule.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



Regions (p) and pi (π) bonds

- Covalent bonds are classified as sigma (σ) and pi (π).
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- Very simple bond is a covalent bond.
- Non-overlapping bond consists of two regions and two pi (π) bonds.



pi bonds are stronger than sigma.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a C-C bond) or partial charges on the C-C bond.
- If the difference in electronegativity is between 0.5 and 1.7, one atom has a partial negative charge and the other has a partial positive charge (like a C-O bond).
- A difference of 1.7 or higher one or both the atoms have a large partial and a small bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

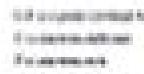


Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Electrons are more attracted towards the more electronegative atom.

- The dipole moment is shown as a vector pointing from the less electronegative towards the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a distance and area.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of individual bond polarities. The dipole moment is measured. A good example is water, dipole.



Sum of the bond dipole moment for the water.

Bond Length and Bond Strength

Reactions of bond length



Length: 1.07 Å, Strength: 1.01 kcal/mol

Length: 1.27 Å, Strength: 0.71 kcal/mol

Length: 1.47 Å, Strength: 0.41 kcal/mol

Length: 1.66 Å, Strength: 0.11 kcal/mol

Length: 1.86 Å, Strength: 0.01 kcal/mol

Length: 2.06 Å, Strength: 0.00 kcal/mol

Length: 2.26 Å, Strength: 0.00 kcal/mol

Length: 2.46 Å, Strength: 0.00 kcal/mol

Length: 2.66 Å, Strength: 0.00 kcal/mol

Length: 2.86 Å, Strength: 0.00 kcal/mol

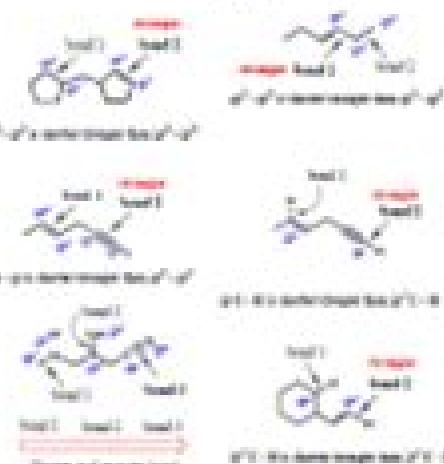
Reactions of bond strength

Average values of C-C and C-H bond depending on the reaction

C-C Bond	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$	$\text{CH}_3\text{CH}_2\text{OH}$
Length	1.37 Å	1.54 Å	1.73 Å
Strength	1.01 kcal/mol	0.71 kcal/mol	0.31 kcal/mol
Hybridization	sp^2	sp^3	sp^3
Efficiency	25.7% 25.7% 25.7%	25.7% 25.7% 25.7%	25.7% 25.7% 25.7%
Length	1.07 Å	1.07 Å	1.07 Å
Strength	0.01 kcal/mol 0.01 kcal/mol 0.01 kcal/mol	0.01 kcal/mol 0.01 kcal/mol 0.01 kcal/mol	0.01 kcal/mol 0.01 kcal/mol 0.01 kcal/mol

Higher the bond length and stronger bond

Example



Properties of covalent bonds

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however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It's just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

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