

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonds

There are two types of bonds: ionic and covalent bonds.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



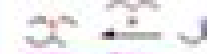
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Difference Between Ionic and Covalent Bonds



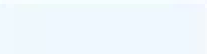
Bond and Electron Sharing



Bonding in Molecules



There are two types of bonds: ionic and covalent bonds.



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the bond is nonpolar.
- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.
- If the difference in electronegativity is greater than 1.7, the bond is ionic.

- A difference of 1.7 or higher indicates that the electrons are not shared equally and a polar bond is formed.

- A difference of 1.7 or higher indicates that the electrons are not shared equally and a polar bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- δ⁺ (delta plus) indicates a partial positive charge.
- δ⁻ (delta minus) indicates a partial negative charge.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The length of a bond is important.

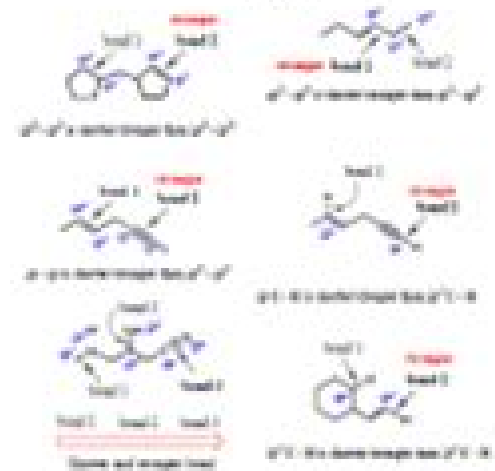


The length of a bond is important.

| Average values of C-C and C-H bonds depending on hybridization | | | |
|--|----------------------------------|----------------------------------|--------------|
| C-C Bond | sp ³ -C-C | sp ² -C-C | sp-C-C |
| Hybridization | sp ³ -sp ³ | sp ² -sp ² | sp-sp |
| Length | 1.54 Å | 1.34 Å | 1.20 Å |
| Strength | 33 kcal/mol | 112 kcal/mol | 160 kcal/mol |
| Hybridization | sp ³ -sp ³ | sp ² -sp ² | sp-sp |
| C-H Bond | sp ³ -H | sp ² -H | sp-H |
| Length | 1.09 Å | 1.08 Å | 1.06 Å |
| Strength | 99 kcal/mol | 110 kcal/mol | 120 kcal/mol |

Higher s character means stronger bond

Examples



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**Herbert Meislich, Howard
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Organic Chemistry T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016 **Study Guide and Solutions**
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Study Guide and Solutions Manual for Essential Organic Chemistry Paula Yurkanis Bruice, 2015-03-04 This Study Guide Solution Manual contains learning objectives chapter summaries and outlines as well as examples self tests and concept questions as well as complete step by step solutions to selected problems **Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us ORGANIC CHEMISTRY STUDY GUIDE Marc Loudon, *Organic Chemistry, Study Guide and Solutions Manual* T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small

however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

A Self-study Guide to the Principles of Organic Chemistry Jiben Roy,2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix

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