

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



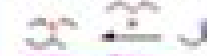
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding

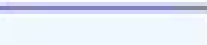


Bonding



There are two types of bonding

Definition of Polarity



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two double bonds** consist of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be polar and nonpolar.

• **Polar covalent bonds** are formed between atoms with different electronegativities.

• **Nonpolar covalent bonds** are formed between atoms with similar electronegativities.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent. Examples are the C-Cl and C-Br bonds.

• If the difference in electronegativity is between 1.7 and 2.0, the bond is ionic. Examples are the Na-Cl and K-Br bonds.

• A difference of 2.0 or higher indicates that the electrons are not shared but are lost to the more electronegative atom.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The vector dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• **δ⁺** (delta plus) indicates an electron-deficient atom.

• **δ⁻** (delta minus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

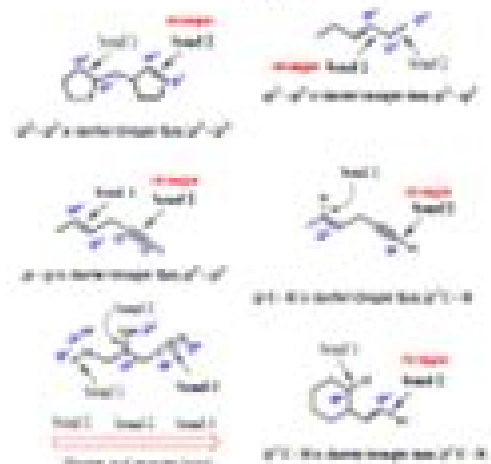


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	460 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	410 kJ/mol	460 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



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