

# Reactions of Amines and Epoxydes

## Preparation of Amines

Ammonium hydroxide:  $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



## Four different types of bonds

1. Ionic bond  
2. Covalent bond  
3. Hydrogen bond  
4. Metallic bond



## The Behavior of the bond



## Strong and Weak bond



## Distance between atoms



## Strength of bonds



## Hydrogen bond



## Electrostatic bond



## Metallic bond



## Covalent bond

## Four different types of bonds

There are four types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges)



## Covalent bond and polar covalent bond

- Covalent bonds are formed by sharing of electron between two atoms
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## Strong bond or longer the bond

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like oxygen atoms in the O<sub>2</sub> and C<sub>6</sub>H<sub>6</sub> bonds).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond between the carbon in nitrogen in N<sub>2</sub>, O<sub>2</sub>, F and other halogens).
- A difference of 1.7 or higher one of both the atoms can form a longer bond and a polar bond is formed.

## A Summary of Ring-opening Reactions of Epoxydes with Strong and Weak Protagonists



### Protagonists and Reactions

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Dipole Moment

The polarity of a bond is denoted by the dipole moment ( $\mu$ ).

1. Nonpolar covalent bond  
2. Polar covalent bond  
3. Ionic bond



- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector represents the size for a partial negative charge and the less electronegative atom, a partial positive charge (O).
- The dipole bond is formed by the difference in electronegativity.
- The dipole bond is formed by the difference in electronegativity.

### Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the dipole moment. The polar nature of dipole moment is measured. A good example is water, dipole.



Water has a net molecular dipole moment of 1.85 D.

## Bond Length and Bond Strength

### Relative bond length



Bond length: H-H > H-Cl > H-Br > H-I

Strength: H-H < H-Cl < H-Br < H-I

Electrostatic bond > Covalent bond > Ionic bond

Length: H-H > H-Cl > H-Br > H-I

Strength: H-H < H-Cl < H-Br < H-I

Ionic bond > Covalent bond > Electrostatic bond

Length: H-H > H-Cl > H-Br > H-I

Strength: H-H < H-Cl < H-Br < H-I

Electrostatic bond > Covalent bond > Ionic bond

### Relative bond strength

## Average values of C-C and C-H bond depending on the bond type

C-C bond:  $3.5 \text{ eV} - \text{C}_2\text{H}_2$ ,  $4.2 \text{ eV} - \text{C}_2\text{H}_4$ ,  $5.0 \text{ eV} - \text{C}_2\text{H}_6$

Hybridization:  $1.7 \text{ eV} - \text{sp}^2$ ,  $1.8 \text{ eV} - \text{sp}^3$ ,  $1.9 \text{ eV} - \text{sp}$

Length:  $1.75 \text{ \AA}$ ,  $1.55 \text{ \AA}$ ,  $1.35 \text{ \AA}$

Strength:  $1.0 \text{ eV} - \text{C}_2\text{H}_2$ ,  $1.1 \text{ eV} - \text{C}_2\text{H}_4$ ,  $1.2 \text{ eV} - \text{C}_2\text{H}_6$

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Hybridization:  $1.7 \$

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**T. W. Graham Solomons**

## Organic Chemistry Final Study Guide:

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*Study Guide with Solutions to Accompany Burgoyne, A Study Course in Organic Chemistry* N. Sue Barnes,1979

**Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry** McMurry,Susan McMurry,1990

*Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry* John McMurry,Susan McMurry,1986

*Study Guide for Organic Chemistry* Seyhan N. Ege,1999

*Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

*PSI Study Guide for Organic Chemistry* by Douglas C. Neckers, Michael P. Doyle E. C. Blossey,1977

*DAT: Dental Admissions Test: Includes 3 Full Length Practice Tests + Online Access to Video Tutorials* Barron's Educational Series,Joseph DiRienzo,John J. Ference,Nicole D. Cornell,Edwin H. Hines,John Swartwood,2018-05-15 This brand new manual prepares dental school applicants across the United States and Canada to pass the required admissions test It features Three full length model tests including a diagnostic test All answers explained in detail Access to video tutorials from the authors and more Test takers will also find thorough reviews of all DAT test topics a general survey of the natural sciences including biology chemistry and

organic chemistry as well as testing for perceptual ability reading comprehension and quantitative reasoning **ONLINE PRACTICE TEST** Students will also get access to one additional full length online DAT test with all questions answered and explained This online exam can be easily accessed by smartphone tablet or computer **Organic Chemistry, Study Guide** T. W. Graham Solomons, 1988-03-08 This class tested text reflects the refinements made in previous editions while expanding updating and adding new information on many important topics Adopting a bio organic emphasis it introduces functional groups early in the text providing an overview of and preparation for subsequent discussions This edition includes increased coverage of Carbon 13 spectra an expanded treatment of conformational effects of molecules with two sections covering basic geometries of molecules and an in depth overview of the subject plus new material on transition metal chemistry and carbohydrate metabolism Worked out examples have been added to chapters on synthesis and end of chapter problems have been expanded by more than 200 The text also includes exceptional full color graphics illustrating conformational effects and general stereochemical properties of organic substances **Class 10 Chemistry MCQ (Multiple Choice Questions)**

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*Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

### **Student Study Guide and Solutions Manual to accompany General**

**Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

*Organic Chemistry* Marye Anne Fox, James K. Whitesell, 2004 Accompanying CD ROM has been enhanced with updated animated illustrations to accompany the presentations and Chem3D files for helpful structure visualization Page 4 of cover

### **Introduction to the study of organic chemistry** John Wade (DSc.), 1898

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