

ORGANIC CHEMISTRY

REACTIONS

FEATURES OF AN ORGANIC REACTION

Mechanism: Describes the overall reaction using a series of simple steps.

Stoichiometry: Calculate reactants and product masses using the balanced equation and molar masses.

Kinetics: The study of the reaction rate and mechanism.

Theoretical yield: Mass of product given by a complete reaction.

Yield: $\frac{\text{actual}}{\text{theoretical}}$ mass (theoretical yield).

Equilibrium: Reaction does not proceed to completion, instead, it reaches a balanced state of forward and reverse reactions.

Major reaction types:

- Acid
- Base
- Oxidation/Reduction
- Condensation
- Substitution (SN1, SN2)
- Ionic Reaction
- Elimination (E1, E2)
- Cationic
- Hydrolysis
- Addition
- Radical reaction

Important named reactions:

- **Halo-Alkane:** from alkyl halide
- **Friedel-Crafts:** alkyl or acyl group
- **Grignard:** alkyl or acyl group
- **Witt-Renbaum, Clemmensen:** reduce ketone to alkane
- **Wittig:** convert aldehyde/ketone to alkene

KINETICS AND REACTION MECHANISM



Transition state (TS): Maximum on the reaction coordinate curve; the least stable intermediate.

Activation Energy (E_a): Energy of the TS relative to the reactants. The change in enthalpy (ΔH) is $-E_a$ for exothermic, $+E_a$ for endothermic.

Hammett-Laurer postulate: The TS is more like the reactant or product that is closer in energy. endothermic TS is like the product, exothermic TS is like the reactant.

Kinetic vs. thermodynamic control: (a) and (b) describe Thermodynamic stability.

• If (a) is large and negative (exergonic), the product formation is likely controlled by "thermodynamic". Large K_{eq} corresponds to a large amount of product, relative to reactant.

• A large E_a may give rise to "kinetic" control; the energy of the TS controls the reaction, instead of the product/reactant thermodynamics.

Solvent effects: A solvent may stabilize an intermediate, decreasing E_a and increasing the rate of the reaction. Charge-complexes are stabilized by polar solvents.

ORGANIC ACID AND BASE

Acid:

- Electrophile: accepts δ^- ions acids
- Proton donor (Brønsted-Lowry acid)
- example: carboxylic acid

Base:

- Electrophile donor (Lewis base)
- Proton acceptor (Brønsted-Lowry base)
- example: amine

Factors enhancing acid strength (HA):

- Medium H-A bond
- Greater electronegativity of "A"
- Inductive effect of substituents on "A" (electron withdrawal enhances transfer)
- More "s" character in hybrid orbital involved in bond to group that is pulled
- Resonance stabilization conjugate base (A⁻)

Factors enhancing base strength:

- Resonance of conjugate acid/bases
- A base is a nucleophile. Electrons effects which shift electron density to the atom with the lone-pair increases base strength.

ALKANE

Properties:

- Hydrocarbon
- Weak intermolecular forces
- Non-polar, general formula C_nH_{2n+2}
- Saturated C-C-C (109°)

Nomenclature:

- Add "-ane" to prefix
- Lowest substituent by position #
- Haloalkane: substitute halide for H

Cycloalkane (C_nH_{2n}):

- Bicyclic: two fused or bridged rings
- $n = 2$: cyclopropane (highly strained)
- $n = 3$: cyclobutane (some flexibility)
- $n = 4$: cyclopentane (slight puckering)
- $n = 5$: cyclohexane: chair = stable conformation
- Axial: less stable, axial position, "perpendicular" to ring. Equatorial position, in ring "plane"
- Cis: two substituents on up position
- Trans: one up and one down



Synthesis:

- Hydrogenate alkene or alkyne (H_2 , Pt catalyst)
- Free-radical reaction of alkane
- Radical halogenation (Cl_2 , Br_2)
- Friedel-Crafts alkylation

Reactions:

- Combustion: alkane + $O_2 \rightarrow CO_2 + H_2O$
- Halogenation to haloalkane (Cl_2/Br_2 , light or heat)

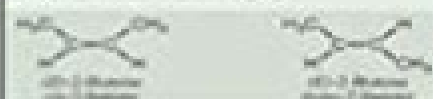
ALKENE $>C=C<$

Properties: Similar to alkane, non-polar, flammable

Nomenclature:

- Add "-ene" to prefix, Use E to denote C=C position
- Saturated C-C-C-C-C-C, unsaturated C=C-C-C
- Polyunsaturated fatty acid: 2 or more C=C
- Alkene: adjacent C-C=C
- Vinyl group: $H_2C=CH-$
- Methylene group: H_2C-
- Allyl group: $H_2C=CH-CH_2-$
- Vinyl halide: halide replaces H on C=C
- Conjugated alkenes: C=C and C=C (consecutive)
- Alkadiene, 2 conjugated C=C, example: butadiene, cis and trans isomers about C=C bonds
- Alkatriene, 3 conjugated C=C
- Aromatic: conjugated monocyclic compound, example: (H) benzene = benzene
- Aromatic: cyclic, non-cyclopentadiene anion, cyclopentadienyl anion (5 electrons)

Isomers: no free rotation of C=C



- E/Z: priority groups by atomic weight of - higher priority groups on the same side
- For monosubstituted: cis is less stable (more branching), the opposite, cis more stable
- **Markovnikov Rule:** from the less substituted alkene
- **Anti-Markovnikov addition:** H adds to C with more H's
- **Zaitsev Elimination:** from alkene with more substitution

Synthesis:

- Dehydration: alcohol (H^+ , heat) (elimination)
- Dehydrohalogenation: haloalkane (base, heat)
- Dehalogenation: vic. dihalide (Zn , acidic aq.)
- Hydrogenation: alkene
- sp^2 , Diels-Alder (H_2 , Pt catalyst)
- sp^2 , Diels-Alder (Cl_2 , SO_2 , $AlCl_3$)
- Wittig: aldehyde/ketone + phosphorus ylide

Reactions:

- Combustion: alkene
- Hydrogen to $2H^+$ alcohol (H^+ , H_2 , Y^- from alkene, see mechanism (Markovnikov))
- Hydrogen to alcohol: hydroboration-oxidation ($2H^+$, H_2 , H_2O_2/CH_3COOH , anti Markovnikov)
- Dehydration: dehydrates to alcohol
- Dehydrohalogenation: alkyl halide (Markovnikov)
- Halogenation (Br_2/H_2O vs. dihaloalkane (N_2O/CH_2 , anti addition)
- Halohydrin (Br_2 , H_2O , anti addition)
- Hydrohalogenation: from a 1,2-diol (alkoxide, acid (H⁺), anti addition)
- Oxidation to carboxylic acid (KMnO₄, hot (H⁺))
- Oxidation to ketone (Os , Zn , H_2)
- Hydrogenation to alkane (H_2 , Pt , anti addition)
- Free-radical polymerization
- Alkylation: Reaction
- allylic halogenation (Cl_2 , heat)
- Diels-Alder: cycloaddition from diene + alkene/diophile

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Maya Shankar Singh, Singh



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Maggon, 2026-01-08 This book is based on the CBSE curriculum and aims to provide students with a clear and concise understanding of key scientific concepts It contains the definitions of various terms that students need to learn to better grasp the subject matter and apply it effectively in their studies By offering simplified explanations this book serves as a valuable tool for quick revision helping students strengthen their foundational knowledge In addition to the theoretical content the book incorporates a variety of illustrations and diagrams to make the learning process more engaging and visually appealing The inclusion of these visuals helps to break down complex ideas into easily digestible chunks and fosters a deeper understanding of the material Each chapter is carefully structured to present the facts and concepts in a straightforward and approachable manner ensuring that even challenging topics are made simple for students to comprehend It is my firm belief that science when taught interactively and engagingly becomes not just a subject to study but a fascinating journey of discovery With this in mind I have tried to create a resource that not only helps students excel academically but also sparks their curiosity and encourages them to explore the world around them I sincerely hope that this book will prove to be a useful companion for students and teachers alike offering both clarity and inspiration Punam Maggon

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Reaction Mechanisms Adam Jacobs,1997-07-17 This book describes the principles that govern chemical reactivity and shows how these principles can be used to make predictions about the mechanisms and outcomes of chemical reactions Molecular orbital theory is used to provide up to date explanations of chemical reactivity in an entirely nonmathematical approach suited to organic chemists A valuable section explains the use of curly arrows vital for describing reaction mechanisms An entire chapter is devoted to exploring the thought processes involved in predicting the mechanisms of unfamiliar reactions Each chapter is followed by a summary of the important points and a selection of problems to help the reader make sure that the material in that chapter has been assimilated The book concludes with a comprehensive glossary of technical terms This text will be of interest to first and second year chemistry undergraduates studying organic chemistry

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Table of Contents Organic Chemistry Reactions Quickstudy Academic

1. Understanding the eBook Organic Chemistry Reactions Quickstudy Academic
 - The Rise of Digital Reading Organic Chemistry Reactions Quickstudy Academic
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Reactions Quickstudy Academic
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Reactions Quickstudy Academic
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Reactions Quickstudy Academic
 - Personalized Recommendations
 - Organic Chemistry Reactions Quickstudy Academic User Reviews and Ratings
 - Organic Chemistry Reactions Quickstudy Academic and Bestseller Lists

5. Accessing Organic Chemistry Reactions Quickstudy Academic Free and Paid eBooks
 - Organic Chemistry Reactions Quickstudy Academic Public Domain eBooks
 - Organic Chemistry Reactions Quickstudy Academic eBook Subscription Services
 - Organic Chemistry Reactions Quickstudy Academic Budget-Friendly Options
6. Navigating Organic Chemistry Reactions Quickstudy Academic eBook Formats
 - ePub, PDF, MOBI, and More
 - Organic Chemistry Reactions Quickstudy Academic Compatibility with Devices
 - Organic Chemistry Reactions Quickstudy Academic Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Reactions Quickstudy Academic
 - Highlighting and Note-Taking Organic Chemistry Reactions Quickstudy Academic
 - Interactive Elements Organic Chemistry Reactions Quickstudy Academic
8. Staying Engaged with Organic Chemistry Reactions Quickstudy Academic
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Reactions Quickstudy Academic
9. Balancing eBooks and Physical Books Organic Chemistry Reactions Quickstudy Academic
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Reactions Quickstudy Academic
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry Reactions Quickstudy Academic
 - Setting Reading Goals Organic Chemistry Reactions Quickstudy Academic
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry Reactions Quickstudy Academic
 - Fact-Checking eBook Content of Organic Chemistry Reactions Quickstudy Academic
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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