

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a bond between chlorine and the C-C bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared in a region of N, O, F and other halogens.
- A difference of 1.7 or higher one or both of the electrons are shared and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

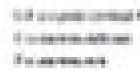


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



Directional dipole moment of the covalent bond.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector (dipole moment) starts from a partial negative charge and the vector (dipole moment) ends in a partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and area.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the individual bond polarities (the dipole moment are canceled). A good example is water, where



Water has a net molecular dipole moment.

Bond Length and Bond Strength

Reactions of bond length



Bond length: 1.07 \AA , 1.27 \AA , 1.47 \AA , 1.66 \AA .

Strength: 10^{12} N/m , 10^{11} N/m , 10^{10} N/m , 10^{9} N/m .

Water is a nonpolar molecule due to the polar and nonpolar bond.

Nonpolar covalent bond: H_2O .



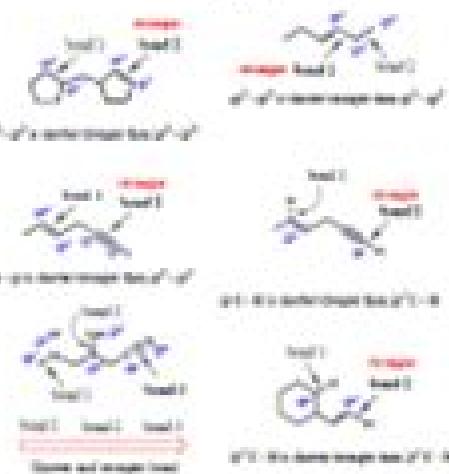
Reactions of bond strength

Average values of C-C and C-H bond depending on the molecule

C-C Bond	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
Length	1.37 \AA	1.54 \AA	1.69 \AA
Strength	10^{12} N/m	10^{11} N/m	10^{10} N/m
Hybridization	sp^2	sp^3	sp^3
Efficiency	25%	30%	35%
Length	1.39 \AA	1.56 \AA	1.71 \AA
Strength	10^{12} N/m	10^{11} N/m	10^{10} N/m

Higher the bond length and stronger bond

Example



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Table of Contents Organic Chem Study Guide

1. Understanding the eBook Organic Chem Study Guide
 - The Rise of Digital Reading Organic Chem Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chem Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chem Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chem Study Guide
 - Personalized Recommendations
 - Organic Chem Study Guide User Reviews and Ratings
 - Organic Chem Study Guide and Bestseller Lists

5. Accessing Organic Chem Study Guide Free and Paid eBooks
 - Organic Chem Study Guide Public Domain eBooks
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 - Organic Chem Study Guide Budget-Friendly Options
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 - Organic Chem Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chem Study Guide
 - Highlighting and Note-Taking Organic Chem Study Guide
 - Interactive Elements Organic Chem Study Guide
8. Staying Engaged with Organic Chem Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chem Study Guide
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chem Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chem Study Guide
 - Setting Reading Goals Organic Chem Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chem Study Guide
 - Fact-Checking eBook Content of Organic Chem Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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