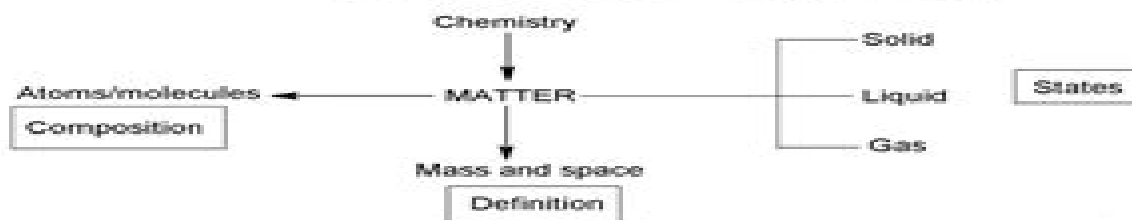


## O level Chemistry Notes



**Atoms:** Building block of matter. It may or may not exist free in nature

Examples: H, He, Na, C

Shells = electrons

Mass = 0

Charge = -1



Nucleus = Protons + Neutrons

Charge = +1                      0

Mass = 1                         1

**Nucleus:** It is central part of an atom.

**Shells:** Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell =  $2n^2$

Where n = no of shell.

1<sup>st</sup> shell = 2es, 2<sup>nd</sup> shell = 8es, 3<sup>rd</sup> shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

**Atomic number:**

Number of protons present in the nucleus of an atom of an element.

Examples:  $_{11}\text{Na}$ ,  $_6\text{C}$ ,  $_8\text{O}$

**Atomic mass or Mass number or nucleon number:** It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples:  $^{23}\text{Na}$ ,  $^{12}\text{C}$ ,  $^{16}\text{O}$

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**N Elkhadragy**



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## **Table of Contents O Level Chemistry Notes**

1. Understanding the eBook O Level Chemistry Notes
  - The Rise of Digital Reading O Level Chemistry Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying O Level Chemistry Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an O Level Chemistry Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from O Level Chemistry Notes
  - Personalized Recommendations
  - O Level Chemistry Notes User Reviews and Ratings
  - O Level Chemistry Notes and Bestseller Lists
5. Accessing O Level Chemistry Notes Free and Paid eBooks
  - O Level Chemistry Notes Public Domain eBooks
  - O Level Chemistry Notes eBook Subscription Services
  - O Level Chemistry Notes Budget-Friendly Options

6. Navigating O Level Chemistry Notes eBook Formats
  - ePub, PDF, MOBI, and More
  - O Level Chemistry Notes Compatibility with Devices
  - O Level Chemistry Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of O Level Chemistry Notes
  - Highlighting and Note-Taking O Level Chemistry Notes
  - Interactive Elements O Level Chemistry Notes
8. Staying Engaged with O Level Chemistry Notes
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers O Level Chemistry Notes
9. Balancing eBooks and Physical Books O Level Chemistry Notes
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection O Level Chemistry Notes
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine O Level Chemistry Notes
  - Setting Reading Goals O Level Chemistry Notes
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of O Level Chemistry Notes
  - Fact-Checking eBook Content of O Level Chemistry Notes
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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