

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Resonance Structures



Resonance structures



There are two types of electronegativity

Electronegativity of elements can be measured by the Pauling scale.



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond is a sigma bond.**
- **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond) or a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- **δ⁻ (delta minus)** indicates an electron deficient atom.
- **δ⁺ (delta plus)** indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

| Bond | H-H | H-Cl | H-Br | H-I |
|----------|------------|------------|------------|------------|
| Length | 109 pm | 127 pm | 141 pm | 154 pm |
| Strength | 436 kJ/mol | 431 kJ/mol | 366 kJ/mol | 298 kJ/mol |

Double bonds are stronger than single bonds and are shorter.

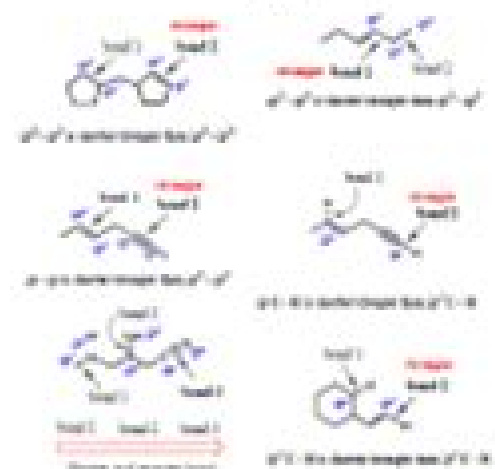


The longer the bond, the weaker it is.

| Average values of C-C and C-H bonds depending on hybridization | | | |
|--|----------------------------------|----------------------------------|----------------------------------|
| C-C Bond | sp ³ -sp ³ | sp ² -sp ² | sp ³ -sp ² |
| Hybridization | sp ³ -sp ³ | sp ² -sp ² | sp ³ -sp ² |
| Length | 154 pm | 134 pm | 147 pm |
| Strength | 376 kJ/mol | 460 kJ/mol | 410 kJ/mol |
| Hybridization | sp ³ -sp ³ | sp ² -sp ² | sp ³ -sp ² |
| C-H Bond | sp ³ -sp ³ | sp ² -sp ² | sp ³ -sp ² |
| Length | 109 pm | 108 pm | 108 pm |
| Strength | 410 kJ/mol | 460 kJ/mol | 410 kJ/mol |

Higher s character means stronger bond

Examples



Organic Chemistry Reactions Study Guide

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Organic 1 and 2 Chemical Reactions Study Guide Milin Kurup, 2020-04-22 This Organic Study Guide and workbook was created by the author in hopes to help students have an organized collection of all reactions studied in Organic 1 and Organic 2 Many students in this class also have hopes of using this study guide as an MCAT review guide this organized set can be useful for students to review before class exams and standardized test This was created by a renowned student Milin Kurup at the University of Florida double majoring in Microbiology and Cognitive and Behavioral Neuroscience As a student of Professor Dr Laura Peterson UF Chemistry Organic Division this guide included all notes a series of 140 chemical reactions and key mechanisms necessary for a comprehensive understanding of the subject

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