

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Types of Bonding

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



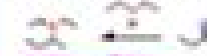
- **Covalent bonds** are formed by sharing of electrons between two atoms.



### The Relative Strength of Ions



### Bond and Bonding



### Bonding



### There are two types of bonding

Definition of ionic and covalent bonds



## Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

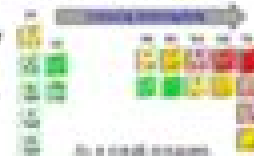
- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are sigma and pi bonds.



pi bonds are stronger than sigma bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the bond is nonpolar.
- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.
- If the difference in electronegativity is greater than 1.7, the bond is ionic.

A difference of 1.7 or higher is a high difference in electronegativity, and a lower difference is a lower difference in electronegativity.

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## Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ<sup>-</sup>) and the less electronegative atom has a partial positive charge (δ<sup>+</sup>).
- δ<sup>-</sup> (delta minus) indicates an electron-deficient atom.
- δ<sup>+</sup> (delta plus) indicates an electron-rich atom.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule, and the dipoles cancel out.

## Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

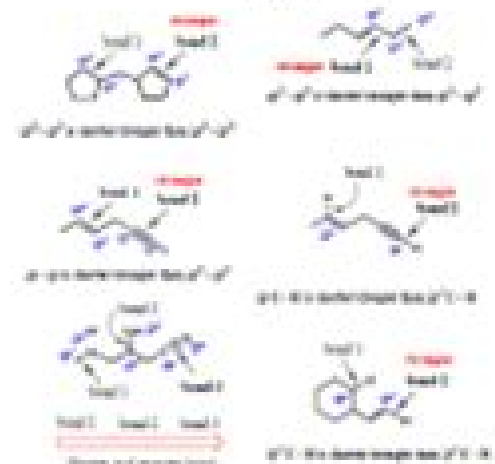


The length of a bond is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp <sup>3</sup> -C-C	sp <sup>2</sup> -C-C	sp-C-C
Hybridization	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp <sup>3</sup> -sp <sup>2</sup>	sp <sup>2</sup> -sp	sp-sp
Length	147 pm	133 pm	121 pm
Strength	356 kJ/mol	473 kJ/mol	591 kJ/mol

Higher s character means stronger bond

## Examples



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