

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for substituted hydroxyl compounds



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two different bond types

There are two types of bonding: **ionic bonding** and **covalent bonding**

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of an



Bonded and Unbonded electron



Electron movement



There are two main categories

1. Definition of electron pair



Sigma (σ) and pi (π) bonds

Covalent bonds are classified in **sigma (σ) bonds** and **pi (π) bonds**

- **Sigma bonds** are formed by head-on (frontal) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a **sigma bond**.
- **One multiple bond** consists of one **sigma** and **n pi bonds**.



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is **between 0.5 and 1.7**, we have a **polar covalent bond**.
- If the difference in electronegativity is **greater than 1.7**, we have an **ionic bond**.

A difference of 1.7 or higher can be high, but the difference can be higher (shared) and a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.



- The **dipole moment** is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- **δ⁻** (delta minus) indicates an electron-deficient atom.
- **δ⁺** (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the **molecular polarity**, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, even if the dipole moments are cancelled. A good example is carbon dioxide.



Cancel the vector or the other pointing point in the same way.

Bond Length and Bond Strength

As the bond length increases, the bond strength decreases.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

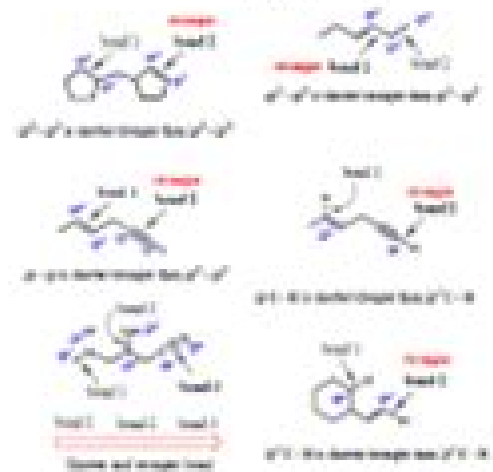


As the bond length increases, the bond strength decreases.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C-C	sp ² -C-C	sp-C-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
C-H Bond	sp ³ -H	sp ² -H	sp-H
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	505 kJ/mol	548 kJ/mol

Higher s character = stronger and shorter bond

Examples



Organic Chem Final Study Guide

Mary H. Bailey



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Table of Contents Organic Chem Final Study Guide

1. Understanding the eBook Organic Chem Final Study Guide
 - The Rise of Digital Reading Organic Chem Final Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chem Final Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chem Final Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chem Final Study Guide
 - Personalized Recommendations
 - Organic Chem Final Study Guide User Reviews and Ratings
 - Organic Chem Final Study Guide and Bestseller Lists

5. Accessing Organic Chem Final Study Guide Free and Paid eBooks
 - Organic Chem Final Study Guide Public Domain eBooks
 - Organic Chem Final Study Guide eBook Subscription Services
 - Organic Chem Final Study Guide Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Organic Chem Final Study Guide Compatibility with Devices
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chem Final Study Guide
 - Highlighting and Note-Taking Organic Chem Final Study Guide
 - Interactive Elements Organic Chem Final Study Guide
8. Staying Engaged with Organic Chem Final Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chem Final Study Guide
9. Balancing eBooks and Physical Books Organic Chem Final Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chem Final Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chem Final Study Guide
 - Setting Reading Goals Organic Chem Final Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chem Final Study Guide
 - Fact-Checking eBook Content of Organic Chem Final Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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