

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding between atoms in a molecule.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



Regions (p) and pi (π) bonds

- Covalent bonds are classified as sigma (σ) and pi (π).
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- Very simple bond is a covalent bond.
- Non-overlapping bond consists of two regions and two pi (π) bonds.



pi bonds are stronger than sigma.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a C-C and C-H bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared in a region of N, O, F and other halogens.
- A difference of 1.7 or higher one of both the atoms can have a larger share and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles

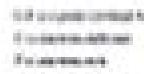


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Directional dipole moment of the covalent bond.

- The dipole moment is shown as a vector pointing from the less electronegative to the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the bond's dipole moment.
- The dipole moment is shown as a horizontal vector.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. The dipole moment of the molecule is measured. A good example is water, dipole.



Sum of bond dipole moment for the water molecule.

Bond Length and Bond Strength

Reactions of bond length



Bond length: 1.07 \AA , 1.27 \AA , 1.47 \AA , 1.67 \AA .

Strength: 100 kJ/mol , 110 kJ/mol , 120 kJ/mol , 130 kJ/mol .

Water molecule has the shortest bond length and strongest bond.

Hydrogen bond: 1.02 \AA , 1.12 \AA .

Hydrogen bond: 100 kJ/mol , 110 kJ/mol .

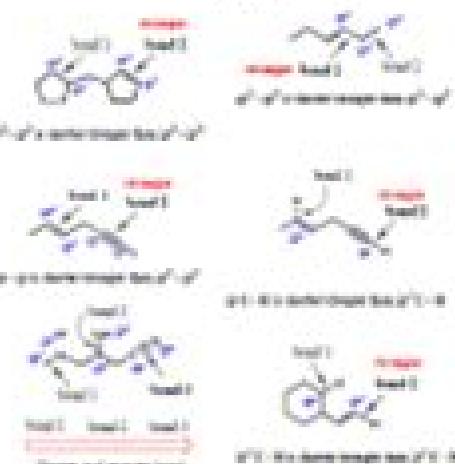
Reactions of bond strength

Average values of C-C and C-H bond depending on the nucleophile

C-C Bond	$\text{H}_2\text{C}-\text{CH}_3$	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$
Hydrogen	100 kJ/mol	100 kJ/mol	100 kJ/mol
Length	1.07 \AA	1.27 \AA	1.47 \AA
Strength	100 kJ/mol	110 kJ/mol	120 kJ/mol
Hydrogen	100 kJ/mol	100 kJ/mol	100 kJ/mol
Hydrogen	100 kJ/mol	100 kJ/mol	100 kJ/mol
Length	1.07 \AA	1.27 \AA	1.47 \AA
Strength	100 kJ/mol	110 kJ/mol	120 kJ/mol

Hydrogen bond and strong bond

Example



Organic Chem Final Study Guide

Mary H. Bailey

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