

## AP Questions: Thermochemistry

1984 B

Substance	Standard Heat of Formation, $\Delta H_f^\circ$ , in kJ/mol
C(s)	0.00
CO <sub>2</sub> (g)	-393.5
H <sub>2</sub> (g)	0.00
H <sub>2</sub> O(l)	-285.85
O <sub>2</sub> (g)	0.00
C <sub>3</sub> H <sub>6</sub> COOH(l)	?

The enthalpy change for the combustion of butyric acid at 25°C,  $\Delta H_{\text{comb}}$ , is -2,183.5 kilojoules per mole. The combustion reaction is

$$\text{C}_3\text{H}_6\text{COOH(l)} + 5 \text{O}_2(\text{g}) \rightarrow 4 \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O(l)}$$

- From the above data, calculate the standard heat of formation,  $\Delta H_f^\circ$ , for butyric acid.
- Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

Substance	Enthalpy of Combustion, $\Delta H$ , in kilojoules/mole
C(s)	-393.5
H <sub>2</sub> (g)	-285.8
C <sub>2</sub> H <sub>5</sub> OH(l)	-1366.7
H <sub>2</sub> O(l)	—

- Write a separate, balanced chemical equation for the combustion of each of the following: C(s), H<sub>2</sub>(g), and C<sub>2</sub>H<sub>5</sub>OH(l). Consider the only products to be CO<sub>2</sub> and/or H<sub>2</sub>O(l).
- In principle, ethanol can be prepared by the following reaction:



Calculate the standard enthalpy change,  $\Delta H$ , for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard molar enthalpy of neutralization of a strong acid by a strong base. Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- What equipment would be needed?
- What measurements should be taken?
- Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, C<sub>3</sub>H<sub>8</sub>, is a hydrocarbon that is commonly used as fuel for cooking.

- Write a balanced equation for the complete combustion of propane gas, which yields CO<sub>2</sub>(g) and H<sub>2</sub>O(l).
- Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O<sub>2</sub> by volume.
- The heat of combustion of propane is -2,220.1 kJ/mol. Calculate the heat of formation,  $\Delta H_f^\circ$ , of propane given that  $\Delta H_f^\circ$  of H<sub>2</sub>O(l) = -285.8 kJ/mol and  $\Delta H_f^\circ$  of CO<sub>2</sub>(g) = -393.5 kJ/mol.
- Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 kilograms of water (specific heat = 4.18 J/g·K), calculate the increase in temperature of water.

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**University of Illinois at Urbana-  
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