

4) Oxidation-Reduction Reactions

Oxidation - is a loss of electrons
 - is an increase in oxidation number

Reduction - is a gain of electrons
 - is a decrease in oxidation number

Oxidizing Agent - brings about the oxidation of another species
 ↗ gets reduced

Reducing Agent - brings about the reduction of another species
 ↗ causes something else to gain electrons by losing
 ↗ gets oxidized

→ OIL RIG (Oxidation Is Loss; Reduction Is Gain)
 LEO~GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal $\xrightarrow{\text{ox}}$ metal ion



2. Metal $\xrightarrow{\text{ox}}$ metallous ion $\xrightarrow{\text{ox}}$ metallic ion



3. Halide ion $\xrightarrow{\text{ox}}$ halogen $\xrightarrow{\text{ox}}$ halite ion $\xrightarrow{\text{ox}}$ halate ion (similar for $\text{NO}_2 \rightarrow \text{NO}_3^-$)



MEMOIZE!

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. $\text{KMnO}_4 \rightarrow \text{Mn}^{2+}$ (acidified solution)



==== Identify the substances being oxidized and reduced in each of the following. ===



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