

## Chapter 12 Review 2

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- 1 A mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together is called a(n)
  - a. dipole.
  - b. Lewis structure.
  - c. chemical bond.
  - d. London force.
- 2 The electrons involved in the formation of a chemical bond are called
  - a. dipoles.
  - b. *s* electrons.
  - c. Lewis electrons.
  - d. valence electrons.
- 3 A chemical bond resulting from the electrostatic attraction between positive and negative ions is called a(n)
  - a. covalent bond.
  - b. ionic bond.
  - c. charged bond.
  - d. dipole bond.
- 4 If two covalently bonded atoms are identical, the bond is
  - a. nonpolar covalent.
  - b. polar covalent.
  - c. nonionic.
  - d. coordinate covalent.
- 5 If the atoms that share electrons have an unequal attraction for the electrons, the bond is called
  - a. nonpolar.
  - b. polar.
  - c. ionic.
  - d. dipolar.
- 6 A covalent bond results when \_\_\_\_\_ are shared.
  - a. ions
  - b. Lewis structures
  - c. electrons
  - d. dipoles
- 7 Nonpolar covalent bonds are not common because
  - a. one atom usually attracts electrons more strongly than the other.
  - b. ions always form when atoms join.
  - c. the electrons usually remain equally distant from both atoms.
  - d. dipoles are rare in nature.
- 8 The greater the electronegativity difference between two bonded atoms, the greater the percentage of
  - a. ionic character.
  - b. covalent character.
  - c. metallic character.
  - d. electron sharing.
- 9 The pair of elements that forms a bond with the least ionic character is
  - a. Na and Cl.
  - b. H and Cl.
  - c. O and Cl.
  - d. Br and Cl.
10. In which of these compounds is the bond between the atoms NOT a nonpolar covalent bond?
  - a.  $\text{Cl}_2$
  - b.  $\text{H}_2$
  - c. HCl
  - d.  $\text{O}_2$

# Modern Chemistry Practice Test Answer Key

**Clemens Wendtner**



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