

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_2\text{O}$
- b.  $\text{K}_2\text{S}$
- c.  $\text{MgCl}_2$
- d.  $\text{CaBr}_2$
- e.  $\text{BaI}_2$
- f.  $\text{Al}_2\text{S}_3$
- g.  $\text{CsBr}$
- h.  $\text{AgF}$

2. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_3\text{N}$
- b.  $\text{K}_2\text{O}$
- c.  $\text{AgBr}$
- d.  $\text{MgI}_2$
- e.  $\text{SrO}$

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide
- b. sodium iodide
- c. silver sulfide
- d. cesium oxide
- d. beryllium iodide
- f. barium hydride
- g. aluminum fluoride
- h. potassium oxide

4. Write the formula for these ionic substances.

- a. silver oxide
- b. aluminum sulfide
- c. sodium nitride
- d. barium chloride
- e. strontium hydride

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**Arthur James Wells**



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### **Table of Contents Modeling Chemistry U8 Ws 3 V2 Key**

1. Understanding the eBook Modeling Chemistry U8 Ws 3 V2 Key
  - The Rise of Digital Reading Modeling Chemistry U8 Ws 3 V2 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U8 Ws 3 V2 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U8 Ws 3 V2 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U8 Ws 3 V2 Key

- Personalized Recommendations
- Modeling Chemistry U8 Ws 3 V2 Key User Reviews and Ratings
- Modeling Chemistry U8 Ws 3 V2 Key and Bestseller Lists
- 5. Accessing Modeling Chemistry U8 Ws 3 V2 Key Free and Paid eBooks
  - Modeling Chemistry U8 Ws 3 V2 Key Public Domain eBooks
  - Modeling Chemistry U8 Ws 3 V2 Key eBook Subscription Services
  - Modeling Chemistry U8 Ws 3 V2 Key Budget-Friendly Options
- 6. Navigating Modeling Chemistry U8 Ws 3 V2 Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry U8 Ws 3 V2 Key Compatibility with Devices
  - Modeling Chemistry U8 Ws 3 V2 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U8 Ws 3 V2 Key
  - Highlighting and Note-Taking Modeling Chemistry U8 Ws 3 V2 Key
  - Interactive Elements Modeling Chemistry U8 Ws 3 V2 Key
- 8. Staying Engaged with Modeling Chemistry U8 Ws 3 V2 Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry U8 Ws 3 V2 Key
- 9. Balancing eBooks and Physical Books Modeling Chemistry U8 Ws 3 V2 Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U8 Ws 3 V2 Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U8 Ws 3 V2 Key
  - Setting Reading Goals Modeling Chemistry U8 Ws 3 V2 Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U8 Ws 3 V2 Key

- Fact-Checking eBook Content of Modeling Chemistry U8 Ws 3 V2 Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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