

Unit 6 – Worksheet 3

Ionic Compounds

Properties

Basic structural unit

1. Give the name of the following simple binary ionic compounds.
 - a. Na_2O
 - b. K_2S
 - c. MgCl_2
 - d. CaBr_2
 - e. BaI_2
 - f. Al_2S_3
 - g. CsBr
 - h. AgF
2. Give the name of the following simple binary ionic compounds.
 - a. Na_3N
 - b. K_2O
 - c. AgBr
 - d. MgI_2
 - e. SrO
3. Write the formula for the following binary ionic compounds.
 - a. lithium bromide
 - b. sodium iodide
 - c. silver sulfide
 - d. cesium oxide
 - d. beryllium iodide
 - f. barium hydride
 - g. aluminum fluoride
 - h. potassium oxide
4. Write the formula for these ionic substances.
 - a. silver oxide
 - b. aluminum sulfide
 - c. sodium nitride
 - d. barium chloride
 - e. strontium hydride

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Arthur James Wells

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