

### Mixed Stoichiometry Problems

Enter only the number value for your answer. **NO UNITS**. If you enter units, it will not grade your responses correctly. **Round to the nearest whole number**. You must **show your t-chart** to earn credit for correct responses.



- How many moles of  $\text{H}_2$  would be required to produce 5.0 moles of water?
- What mass of  $\text{H}_2\text{O}$  is formed when  $\text{H}_2$  reacts with 384 g of  $\text{O}_2$ ?
- How many moles are there in 67.2 grams of  $\text{H}_2$ ?



- If 30.0 g of  $\text{H}_2\text{SO}_4$  reacts with sodium hydroxide, what mass of water is produced?
- What mass of  $\text{H}_2\text{SO}_4$  would be required to react with 0.75 mol of  $\text{NaOH}$ ?
- Determine the theoretical yield of water when 11.2 g of  $\text{H}_2\text{SO}_4$  reacts with 7.1 g of  $\text{NaOH}$ .



- How many grams of  $\text{SnF}_2$  are produced from the reaction of 30.0 g of  $\text{HF}$ ?
- How many moles of  $\text{HF}$  will be needed to react with 48 grams of Tin?
- How many moles are in  $4.12 \times 10^{24}$  atoms of  $\text{HF}$ ?

4. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide,  $\text{LiOH}$ , according to the following chemical equation.



- If 20.0 moles of  $\text{CO}_2$  is exhaled, the average amount exhaled by a person each day, how much (in grams)  $\text{Li}_2\text{CO}_3(\text{s})$  is produced?

# Mixed Stoichiometry Practice

**British Nuclear Energy Society**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the text 'British Nuclear Energy Society'.

## **Mixed Stoichiometry Practice:**

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