

REDOX REACTIONS.

NOTES...

Intro

Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$

- Sub-points:

- Loss of hydrogen
- Gain of oxygen
- Gain of **electronegative** element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$

- Sub-points:

- Gain of hydrogen
- Loss of oxygen
- Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$

Oxidation Reduction Reactions Study Guide

**Katherine J. Denniston, Denniston
Katherine**

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Table of Contents Oxidation Reduction Reactions Study Guide

1. Understanding the eBook Oxidation Reduction Reactions Study Guide
 - The Rise of Digital Reading Oxidation Reduction Reactions Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Oxidation Reduction Reactions Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Oxidation Reduction Reactions Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Oxidation Reduction Reactions Study Guide
 - Personalized Recommendations
 - Oxidation Reduction Reactions Study Guide User Reviews and Ratings
 - Oxidation Reduction Reactions Study Guide and Bestseller Lists
5. Accessing Oxidation Reduction Reactions Study Guide Free and Paid eBooks

- Oxidation Reduction Reactions Study Guide Public Domain eBooks
- Oxidation Reduction Reactions Study Guide eBook Subscription Services
- Oxidation Reduction Reactions Study Guide Budget-Friendly Options

6. Navigating Oxidation Reduction Reactions Study Guide eBook Formats

- ePUB, PDF, MOBI, and More
- Oxidation Reduction Reactions Study Guide Compatibility with Devices
- Oxidation Reduction Reactions Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Oxidation Reduction Reactions Study Guide
- Highlighting and Note-Taking Oxidation Reduction Reactions Study Guide
- Interactive Elements Oxidation Reduction Reactions Study Guide

8. Staying Engaged with Oxidation Reduction Reactions Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Oxidation Reduction Reactions Study Guide

9. Balancing eBooks and Physical Books Oxidation Reduction Reactions Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Oxidation Reduction Reactions Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Oxidation Reduction Reactions Study Guide

- Setting Reading Goals Oxidation Reduction Reactions Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Oxidation Reduction Reactions Study Guide

- Fact-Checking eBook Content of Oxidation Reduction Reactions Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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