

## Chemical energetics and chemical equilibria

1. Iodine reacts with chlorine to form dark brown iodine monochloride.



This reacts with more chlorine to give yellow iodine trichloride. An equilibrium forms between these iodine chlorides.

(a) What do you understand by the term equilibrium? [2]

(b) When the equilibrium mixture is heated, it becomes a darker brown color.

Suggest if the reverse reaction is endothermic or exothermic.

Give a reason for your choice. [1]

(c) The pressure on the equilibrium mixture is decreased.

(i) How would this affect the position of equilibrium? Give a reason for your choice. It would move to the reason [1]

(ii) Describe what you would observe. [1]

(d) Calculate the overall energy change for the reaction between iodine and chlorine using the bond energy values shown.



Bond	Energy / kJ per mol
I—I	151
Cl—Cl	242
I—Cl	208

Show your working. [3]

(e) Draw a labelled energy level diagram for the reaction between iodine and chlorine using the information in (d)

2. Ammonia is made by the Haber process



(a) State one major use of ammonia. [1]

# Pdf Questions With Answers In Energetics

**Michael Rajnik**

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